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LIMNOLOGY OF SELECTED SOUTH DAKOTA LAKES

BY

ARTWIN E. SCHMIDT

A thesis submitted  
in partial fulfillment of the requirement for  
the degree Master of Science, Major in  
Wildlife Biology, South Dakota  
State University

1968

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## LIMNOLOGY OF SELECTED SOUTH DAKOTA LAKES

This thesis is approved as a creditable and independent investigation by a candidate for the degree, Master of Science, and is acceptable as meeting the thesis requirements for this degree, but without implying that the conclusions reached by the candidate are necessarily the conclusions of the major department.



Thesis Adviser

Date

Wildlife Management Department

# LIMNOLOGY OF SELECTED SOUTH DAKOTA LAKES

## Abstract

ARTWIN E. SCHMIDT

Physical, chemical, and biological aspects of 45 lakes in South Dakota were studied from July 1965 to July 1967. Maximum water temperature at the surface reached 28 C. Most of the lakes studied exhibited continuous circulation except when ice covered. Thermocline formation was observed in six of the lakes. Light transmission was influenced by turbidity, and varied greatly within individual lakes and among lakes. Dissolved oxygen concentrations ranged from near saturation to less than the recommended minimum for fish life. All lakes studied were basic ranging from a pH of 7.1 to 11.3. Specific conductance of lakes occupying open basins was lowest in the unglaciated area west of the Missouri River (70-590 micromhos at 25 C) and highest in Mankato drift of the Wisconsin ice age (330-1260 micromhos at 25 C). Concentrations of major anions and cations tended to follow patterns which were associated with major physical divisions of the state or various drift types of the Wisconsin ice age. Trace elements were found in most lakes studied. Those lakes which developed dense summer blooms of phytoplankton were usually dominated by the blue-green algae Aphanizomenon or Microcystis. The maximum concentration of Chlorophyll a observed was 19.1 mg/l. Chlorophyll concentrations were generally higher in lakes east of the Missouri River.

## ACKNOWLEDGEMENTS

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## INTRODUCTION

An understanding of limnological relationships including physical, chemical, and biological aspects is essential for maximum utilization of water resources. Physical and chemical parameters should be determined first to provide a basis for biotic evaluations. Physical factors are important as they limit length of growing season and affect the chemical and biological constituents. Hayes (1964) found that area and depth of lakes affect standing crop. The level of total dissolved solids is important in determining lake productivity (Northcote and Larkin, 1956). Ball (1948), Carlander (1955), and Turner (1960) reported that standing crop of fishes are directly related to levels of alkalinity in lakes and ponds. Moyle (1956) showed that the type of biological community is influenced by concentrations of chemical nutrients.

Limnological relationships have been intensively studied in other areas, but such information is limited for South Dakota waters. A comparative study of selected physical, chemical, and biological aspects of 45 lakes in South Dakota was initiated in July 1965 and terminated July 1967. The influences of climatic conditions, soil groups, geologic associations, and lake morphometry on biological, physical, and especially chemical factors were considered.

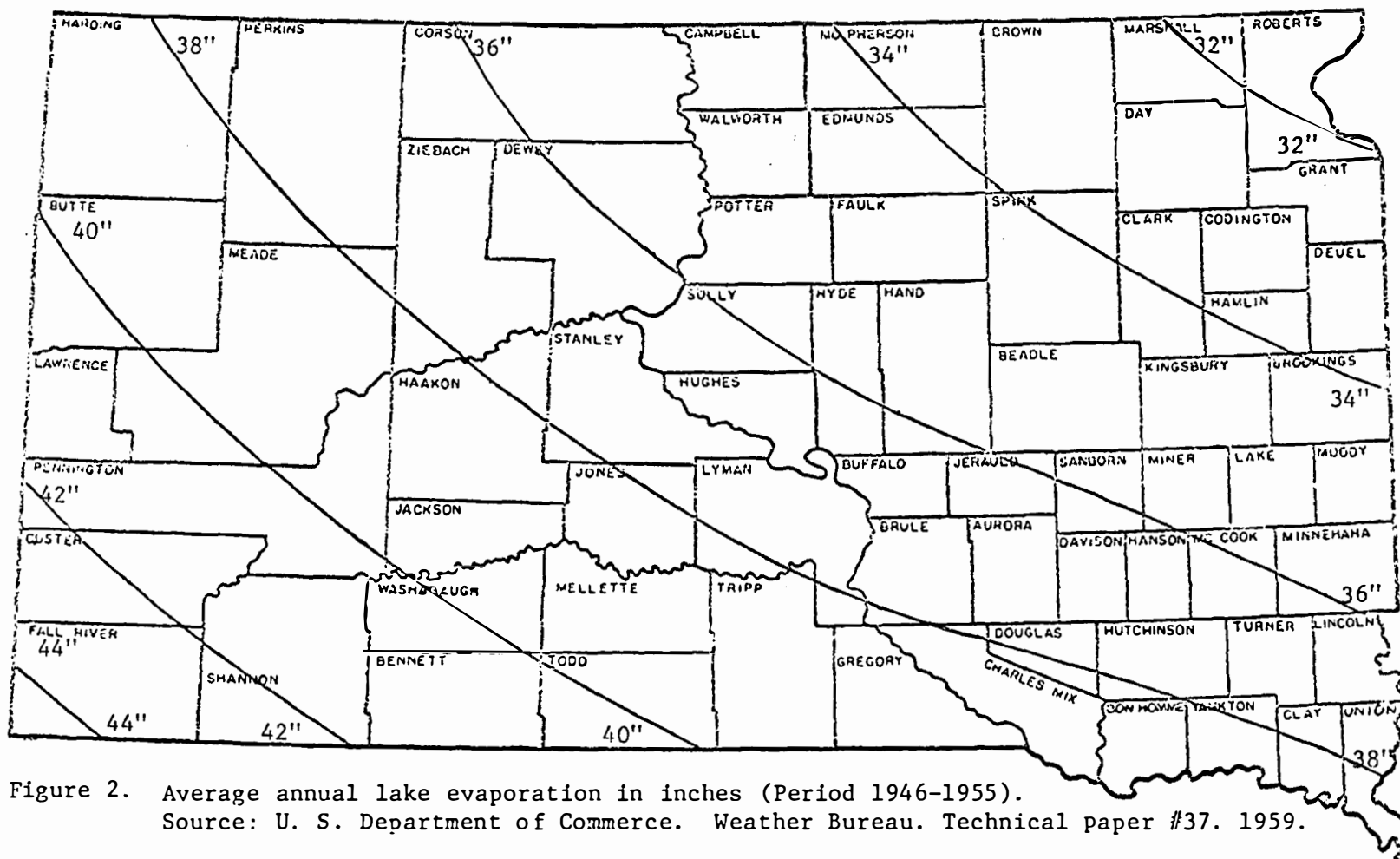
## STUDY AREA

South Dakota has a continental climate with extremes of summer heat, winter cold, and extreme temperature fluctuations. The average number of days without frost varies from 130 days in the northern part of the state to 160 days in the southeastern part (Westin, Puhr, and Buntley, 1959). Annual precipitation varies from 14 inches in the northwest to 24 inches in the southeast, and average temperature varies from 43 to 45 F (Figure 1). Average annual lake evaporation varies from 32 to 44 inches from northeast to southwest (Figure 2). Climatic classification ranges from Moist subhumid in the southeast to Semiarid in the west (Thorntwaite, 1948).

The state can be divided into two main areas; the Central Lowland which occupies the eastern one third, and the Missouri Plateau (Figure 3). The Central Lowland and that portion of the Missouri Plateau east of the Missouri River were glaciated during the Wisconsin glacial age. The Missouri Plateau west of the Missouri River is unglaciated except for localized areas near the river. Soils of the Central Lowland belong to the Chernozem group, while the Missouri Plateau has mainly Chestnut soils.

The Central Lowland and Missouri Plateau were divided into 12 physical divisions. These physical divisions and the natural land forms were described by Flint (1955) and Rothrock (1943). Soil groups within these divisions were described by Westin, et al. (1959).





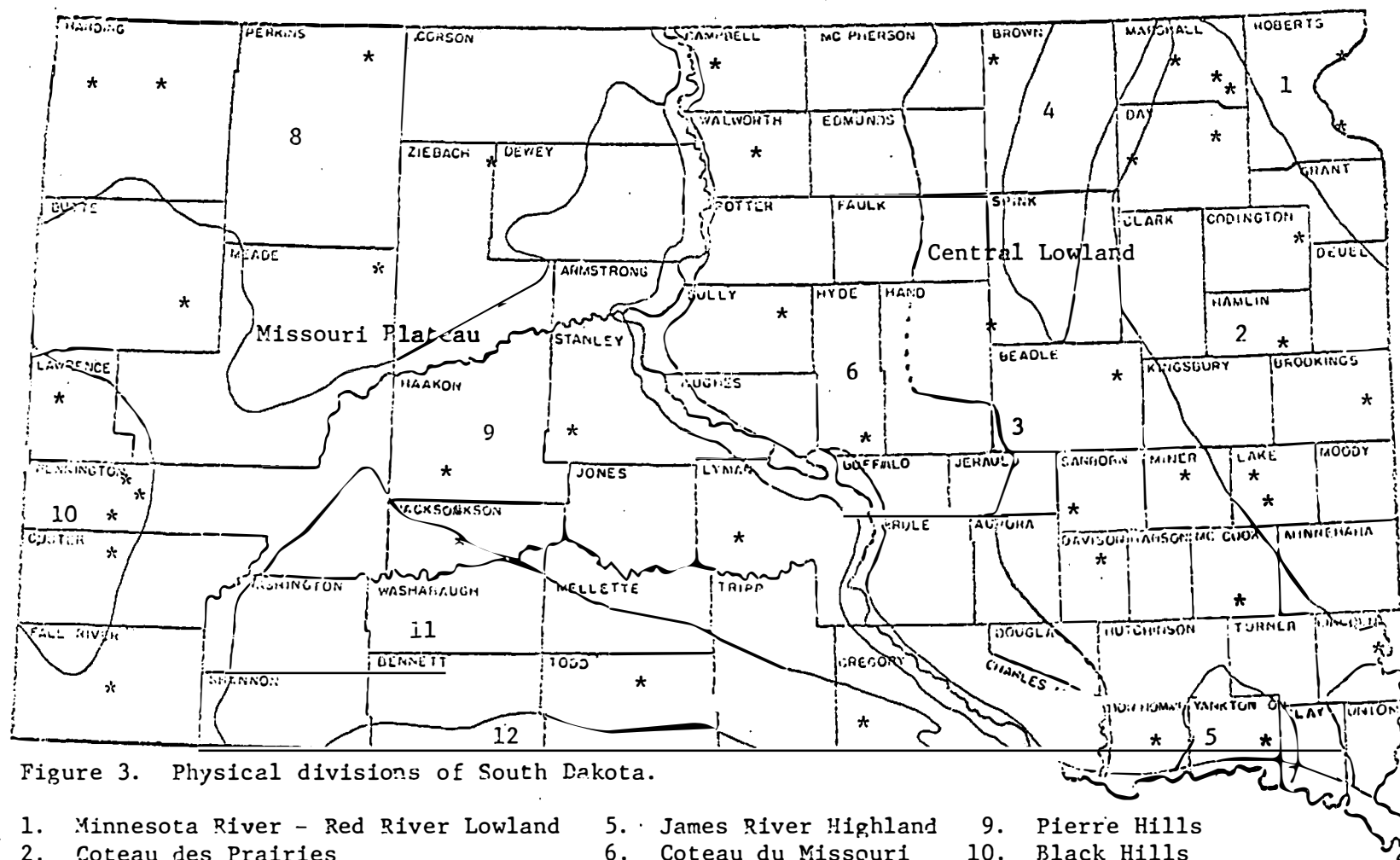


Figure 3. Physical divisions of South Dakota.

- |  |                          |                       |
|--|--------------------------|-----------------------|
| 1. Minnesota River - Red River Lowland | 5. James River Highland  | 9. Pierre Hills       |
| 2. Coteau des Prairies                 | 6. Coteau du Missouri    | 10. Black Hills       |
| 3. James River Lowland                 | 7. Missouri River Trench | 11. Southern Plateaus |
| 4. Lake Dakota Plain                   | 8. Northern Plateaus     | 12. Sand Hills        |

Source: Flint, R. F. Prof. Paper 262, USGS.

\* Location of lakes studied.

Locations of the 45 lakes studied and their positions in physical divisions of South Dakota are shown in Figure 3; morphometric data are given in Table 1.

### Physical Divisions

#### Minnesota River - Red River Lowland

Located in extreme northeastern South Dakota, this division is a broad valley-like area 900 to 1,100 feet above sea level. The continental divide between drainage to the Arctic Ocean and to the Gulf of Mexico is midway between Lake Traverse and Bigstone Lake. The trench occupied by these lakes was formed by overflow from glacial Lake Agassiz. Lakes studied in this division include Bigstone and Traverse.

#### Coteau des Prairies

This division is a plateau irregularly covered with glacial drift. Its rough surface is attributed to morainic deposits of the Wisconsin ice age. The coteau slopes gently to the south where it is drained by the Big Sioux river. Elevations range from 1,600 to 2,000 feet above sea level. Lakes studied in this area were; Buffalo, Enemy Swim, Hendricks, Herman, Madison, Poinsett, Punished Woman, Roy, and White.

The Coteau des Prairies and the Minnesota River - Red River Lowland lie in Chernozem soil which has developed in a cool moist climate. Normal annual precipitation is 20 to 22 inches, while

Table 1. Morphometry of lakes studied.

Lake	County	Type	Surface Area (acres)	Depth in Feet	
				Maximum	Average
Alvin	Lincoln	Artificial	100	30	12
Amsden	Day	Artificial	240	27	9
Angostura	Fall River	Artificial	5,000	79	30
Bigstone	Roberts	Natural	21,120	15	10
Bismarck	Custer	Artificial	22	24	16
Brakke	Lyman	Artificial	135	18	8
Buffalo	Marshall	Natural	1,780	15	6
Burke	Gregory	Artificial	27	16	8
Byron	Beadle	Natural	1,300	10	8
Carthage	Miner	Artificial	38	14	8
Catron Pond (sw 1/4 Sec 19, T18N, R2E)	Harding	Artificial	30	14	6
Cottonwood	Sully	Artificial	1,000	11	6
Durkee	Meade	Artificial	80	16	6
Elm	Brown	Artificial	1,187	35	18
Enemy Swim	Day	Natural	1,975	26	10
Gardner	Harding	Artificial	210	20	10
Hayes	Stanley	Artificial	100	23	10
Henry	Bon Homme	Artificial	70	30	15
Hendricks	Brookings	Natural	1,600	10	5
Herman	Lake	Natural	1,350	8	4
Hiddenwood	Walworth	Artificial	30	15	5
Horse Thief	Pennington	Artificial	19	64	24
Iron Creek	Lawrence	Artificial	23	32	15
Isabell	Dewey	Artificial	130	20	8
Kadoka	Jackson	Artificial	120	14	6
Madison	Lake	Natural	3,200	17	8
Merindahl	Yankton	Artificial	110	30	13
Mission	Todd	Artificial	62	12	6

Table 1. (continued)

Lake	County	Type	Surface Area (acres)	Depth in Feet	
				Maximum	Average
Mitchell	Davison	Artificial	670	29	15
Newell	Butte	Artificial	164	35	13
Pactola	Pennington	Artificial	797	150	76
Pocasse	Campbell	Artificial		30	
Poinsett	Hamlin	Natural	8,000	18	6
Punished Woman	Codington	Natural	400	8	5
Rose Hill	Hand	Artificial	75	32	10
Roy	Marshall	Artificial	200	20	8
Shadehill	Perkins	Artificial	5,000	90	45
Sheridan	Pennington	Artificial	384	90	41
Stephen Mission Dam	Hyde	Artificial	120	18	8
Traverse	Roberts	Natural	11,000	12	9
Twin	Sanborn	Natural	300	12	6
Twin	Spink	Natural	1,500	17	6
Vermillion	McCook	Artificial	550	23	12
Waggoner	Haakon	Artificial	107	30	14
White	Marshall	Artificial	200	20	8

Source: South Dakota Department of Game, Fish, and Parks records.

average temperature is 43 to 45 F. Average annual lake evaporation is 30 to 34 inches. These conditions favor accumulation of organic matter and retard its destruction. Soils of this region are differentiated from soils which adjoin them on the west by having 1) deeper lying horizons of carbonate accumulation, 2) deeper lying horizons of salt accumulation, and 3) higher content of organic matter and total nitrogen in the surface horizons (Westin, et al., 1959).

### James River Lowland

The James River Lowland is a gently undulating plain lying lower than the Coteau des Prairies or the Coteau du Missouri with elevations from 1,300 to 1,400 feet. The area is drained to the south by the James River. Water falling on most of the basin never reaches the stream valley because it is retained in glacier-built hollows until it evaporates or sinks into the ground.

Climatic conditions in the James River Lowland vary. Normal annual precipitation in the north is 18 to 20 inches while average annual temperature varies from 43 to 45 F. Lakes studied in this northern area: Amsden, Byron, Carthage, Elm, Mitchell, Twin Lake in Spink County, and Twin Lake in Sanborn County lie in Chernozem soils which developed in a drier, cooler environment than is found in the southern James River Lowland. The southern Lowland receives 20 to 22 inches normal annual precipitation, and average temperature is 45 to 48 F. These warmer and more moist conditions result in a higher rate

of organic matter destruction and nitrogen release. The depth of carbonate leaching in this region is greater than in any of the other regions. Lakes Alvin, Henry, and Vermillion were in this warm, moist area of the Lowland.

#### James River Highland

The area south of the Lowland is the James River Highland. This division is composed of three ridges of drift covered bedrock. These ridges: Turkey, James, and Yankton are underlain by relatively resistant chalk and limestone. Climatic conditions are similar to those found in the southern James River Lowland. Lake Merindahl was studied in this division.

#### Coteau du Missouri

The portion of the Missouri Plateau lying east of the Missouri River is the Coteau du Missouri. This highland area is covered with glacial drift and underlain by Pierre shale and other formations. Elevations range to heights of over 2,100 feet above sea level. Drainage of the area is eastward to the James River basin and westward to the Missouri River basin. All lakes studied in this area lie in Chestnut soils which are characterized by lighter color and less leaching of carbonates than Chernozem soils. Lakes studied in this division include Cottonwood, Hiddenwood, Pocasse, Rose Hill, and Stephen Mission.

### Missouri Trench

The Missouri Trench averages slightly over a mile in width with the valley floor 300 to 600 feet below the tops of the dissected bluffs. A series of man-made reservoirs occupies the trench. These reservoirs were not studied.

### Missouri Plateau

The Missouri Plateau is the main plateau region from which the Coteau du Missouri was separated by the Missouri River. This plateau occupies the remainder of the state except for the Black Hills and the Sand Hills. Elevations range from approximately 2,000 feet near the Missouri River to over 3,000 feet in the west. All major drainage is to the east. There are three major divisions of the Missouri Plateau in South Dakota.

### Southern Plateaus

The southern division of the Missouri Plateau is the Southern Plateaus. Lakes Burke and Mission were studied in this area. These lakes lie in an extension of the Chernozem soils. Average annual precipitation is 18 to 20 inches, while the mean temperature varies from 45 to 48 F.

### Pierre Hills

The central portion of the Missouri Plateau is the Pierre Hills. It consists of a series of smooth hills and ridges with rounded tops. This area is underlain almost entirely by Pierre shale which breaks

down into a dark pliable clay. The Bad, Cheyenne, and White rivers drain the area eastward to the Missouri River. Lakes studied in the Pierre Hills include Brakke, Hayes, Kadoka, Newell, and Waggoner.

### Northern Plateaus

The Northern Plateaus is a series of plateaus and isolated buttes underlain by Fox Hills sandstone and younger Cretaceous strata. The area is drained eastward by the Grand and Moreau Rivers. The Little Missouri river crosses the northwestern corner of the area, and drains northward. Catron Pond, Durkee, Gardner, Isabell, and Shadehill lakes were studied in this division.

Chestnut soils of the Northern Plateaus and the Pierre Hills receive 14 to 18 inches of moisture annually. Mean temperatures of this area range from 43 to 48 F.

### Black Hills

The Black Hills is a mountainous region formed simultaneously with the Rocky Mountains. It consists of four general areas: 1) a central core of metamorphic and igneous rocks encircled by 2) rings of sedimentary rocks which form high plateaus on the north and west and steep ridges on the east and south, 3) the Red Valley, and 4) the Hogback Ridge. Grey Wooded soils occur in local alluvial positions in the central core and in places on the broad limestone plateaus. These soils, developed in a humid climate with annual precipitation of 20 to 25 inches and an average temperature of 40 to 45 F, have deep

carbonate leaching. Chestnut soils occur in the Red Valley. Lakes studied in the Black Hills were Angostura, Bismarck, Horse Thief, Iron Creek, Pactola, and Sheridan.

## METHODS

Measurements of selected physical parameters and water samples for laboratory analysis were obtained quarterly. Each lake was studied for a one-year period. Lakes in the northern half of South Dakota were sampled from July 1965 to July 1966, and lakes in the southern area were sampled from July 1966 to July 1967. Samples were taken at one to five stations on each lake; the number of stations was determined by lake size. Water samples were taken at 1.5 meter depth intervals in shallow lakes and at the surface, mid-depth, and bottom of deeper lakes. A composite sample from each depth interval was analyzed in the laboratory since time precluded analysis of individual samples. Composite samples were frozen in plastic bags until laboratory analyses were conducted. Analyses of alkalinity, dissolved oxygen concentration, pH, light penetration, and temperature patterns were conducted in the field. Temperature was measured with an electrical resistance bridge thermometer. Light transmission was measured with a relative irradiance meter powered by two photocells in circuit with a galvanometer and variable potentiometer.

Except as otherwise noted chemical analyses followed procedures outlined by Hach Chemical Co. (Hach Chemical Co., Cat. No. 9). Volumetric titrations were used for hardness, alkalinity, sulfate, and chloride. Other Hach procedures utilized were colorimetric. Color development was measured with a Bausch & Lomb "Spectronic 20" spectrophotometer. The mercuric nitrate method was used for chloride

determinations. Sulfate concentrations were determined by titrating with barium perchlorate using Thorin as an indicator (Fritz, 1955). Analysis for dissolved oxygen, phosphate, sodium, and potassium were conducted according to Standard Methods for the Examination of Water and Wastewater (Am. Pub. Health Assoc., 1965). Dissolved oxygen concentrations were determined by the Winkler method. The stannous chloride method for phosphate determination was used only from July 1966 to July 1967. Sodium and potassium concentrations were determined by flame photometry.

Samples for determination of chlorophyll concentrations were obtained in the field by filtering one liter of the composite sample through a cellulose Millipore filter having a pore size of 0.47 microns. Pigments were extracted with 90 per cent acetone according to procedures outlined by the National Research Council (1964). Concentrations of pigments were determined by the method described by Richards with Thompson (1952).

## RESULTS AND DISCUSSION

### Physical Considerations

#### Temperature

The lakes studied were second or third order temperate lakes as defined in Welch (1952). Temperature of the bottom waters of second order lakes varies but not far from 4 C. These lakes have two circulation periods each year. Third order lakes have the temperature of the bottom waters very similar to that of surface water, and exhibit continuous circulation except when ice covered.

Most of the lakes studied were third order lakes since prevailing winds maintained continuous circulation while the water was open. Third order lakes had small temperature variations from top to bottom. This range seldom exceeded 3 C at any given time.

The vertical temperature distributions of second order lakes included in this study are shown in Figure 4. These lakes all exhibited thermocline formation and low hypolimnetic temperatures. Depth to the top of thermocline varied from three to ten meters in separate lakes.

The maximum surface temperature recorded during the study was 28 C. Temperatures of 23 to 25 C were common throughout the summer months. Most lakes studied developed inverse temperature gradients under ice cover. These inverse gradients ranged from 1 C at the surface to about 4 C at the bottom. Inverse temperature gradients may have occurred in all lakes studied but were not observed since sampling was limited to once per season.

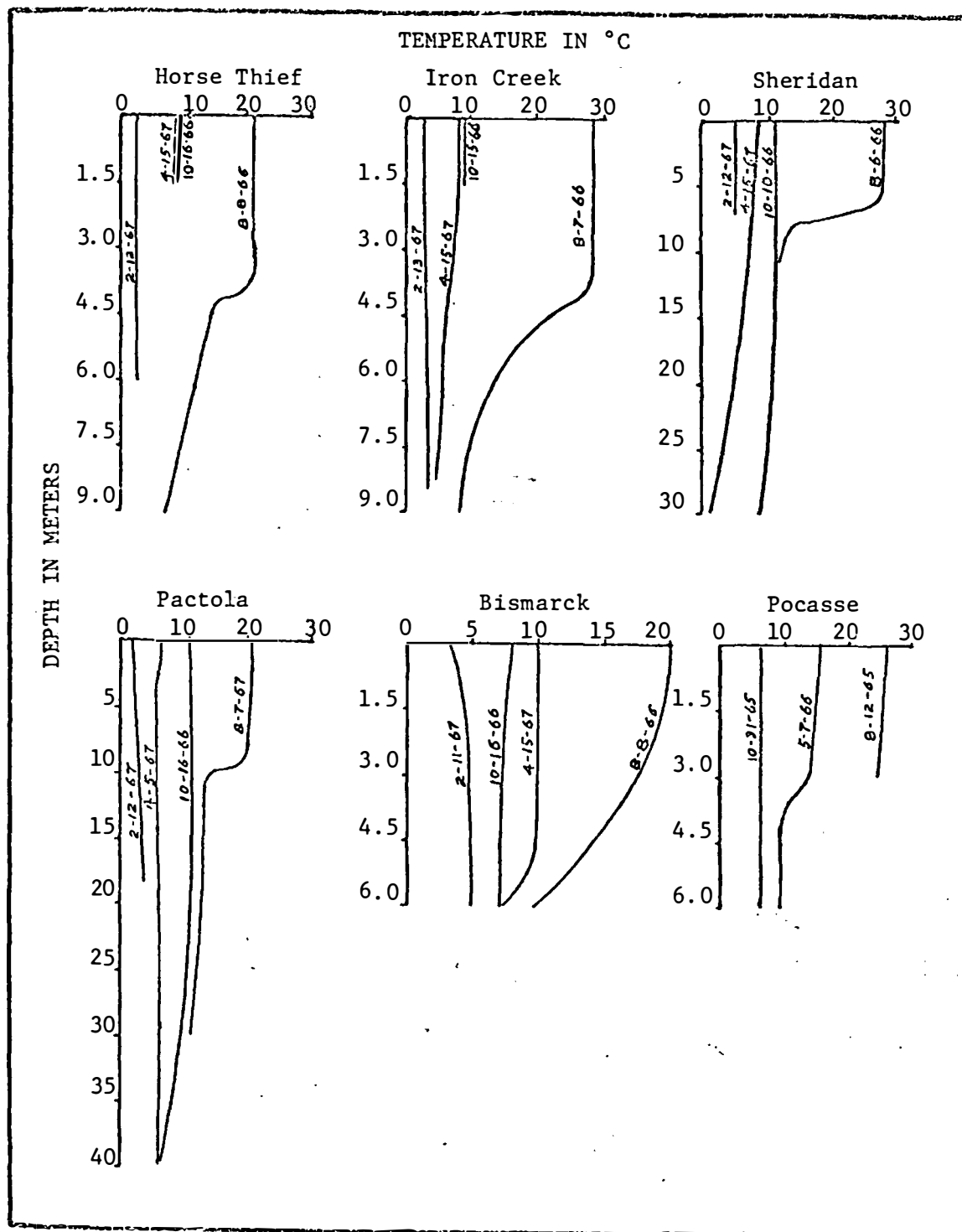


Figure 4. Vertical temperature distributions of second order lakes, by season, July 1965 to July 1967.

Periods of complete ice cover are considerably longer in the northern than in the southern parts of the state. Ice formation in the northeast occurred in mid-November 1965, and break up occurred in mid-April 1966. Ice formed on lakes in the southern portions of the state in mid-November 1966, but ice break up was complete by April 1, 1967. Maximum ice thickness encountered ranged from over 100 centimeters in the northeast to about 20 centimeters in the southwest.

### Light Transmission and Turbidity

Light available for photosynthesis affects the entire aquatic ecosystem through its influence on primary productivity. Available light is dependent upon surface light intensity and subsequent extinction during penetration of water. In general, as water depth increases arithmetically the available light decreases geometrically (Welch, 1952). Many influences modify this relationship, but this rule usually held true for lakes studied. Light transmission values were determined during periods of open water.

Turbidity is the opaqueness of water resulting from the presence of suspended materials. These materials may be settling or non-settling and their sources are innumerable. Among the most important of these are detritus, fine sand, particles of clay, and plankton. Turbidity measurements (Jackson Turbidity Units) are expressed as averages of all depths sampled.

I believe the turbidity readings are biased since all samples were frozen and taken to the laboratory before turbidity was determined. As a check on the reliability of this method turbidity determinations were made on one set of samples before and after freezing for one month at -15 C. After freezing turbidity measurements were 31 per cent higher than before freezing.

Data from series of light transmission and turbidity measurements for lakes studied are shown in Table 2. Light transmission is expressed as relative intensity of light falling on the surface. It is apparent from these data that light transmission varies greatly by season within individual lakes. In general, light transmission shows an inverse relationship to the degree of turbidity. Turbidity was generally lowest in the winter and highest in the summer and fall. This is probably due to stirring of bottom sediments by wind action and the development of dense algal blooms in some lakes in the late summer and fall.

Hutchinson (1957) states that some investigators have concluded that it is better to neglect supposed variations in vertical illumination due to the height of the sun. They claim that the observed values of the slope of the line obtained by plotting intensity logarithmically against depth on semilogarithmic paper do not vary with the angular height of the sun.

The relation between light transmission and depth for some of the lakes studied is shown in Figure 5. This shows a wide variability in light transmission in lakes throughout the state; the depth to

Table 2. Summary of data pertaining to turbidity and light transmission. Light transmission is expressed as relative intensity of surface radiation.

Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at      Depth of 1% 4 cm.      transmission (in meters)		Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at      Depth of 1% 4 cm.      transmission (in meters)	
Alvin				Bismarck			
Aug. 15, 1966	59	50	0.9	Aug. 8, 1966	36	50	4.1
Nov. 5, 1966	48	60	2.1	Oct. 16, 1966	47	70	2.2
Feb. 8, 1967	23			Feb. 11, 1967	43		
Apr. 5, 1967	52	50	1.9	Apr. 15, 1967	39	54	1.4
Amsden				Brakke			
Aug. 6, 1965	24	70	3.5	Aug. 10, 1966	61	45	1.8
Nov. 7, 1965	20	71	3.9	Oct. 8, 1966	45	61	2.4
Feb. 19, 1966	11			Feb. 9, 1967	26		
May 6, 1966		65	1.1	Apr. 12, 1967	49	45	1.0
Angostura				Buffalo			
Aug. 9, 1966	23	66	5.3	July 30, 1965	61	73	2.8
Oct. 17, 1966	20	70	6.6	Nov. 26, 1965	30		
Feb. 11, 1967	13			Feb. 11, 1966	30		
Apr. 16, 1967	22	54	7.5	Apr. 24, 1966	26	65	3.7
Bigstone				Burke			
July 14, 1965	50	35	1.0	Aug. 11, 1966		57	1.5
Nov. 11, 1965	54	37	3.9	Oct. 9, 1966	45	62	2.7
Feb. 18, 1966	65			Feb. 9, 1967	60		
Apr. 23, 1966		74	2.9	Apr. 11, 1967	26	50	4.4

Table 2. (continued)

Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u>		Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u>	
		% at	Depth of 1%			% at	Depth of 1%
		4 cm. transmission	(in meters)			4 cm. transmission	(in meters)
Byron				Durkee			
Aug. 8, 1965	25	42	1.1	Aug. 28, 1965	35	60	2.1
Nov. 5, 1965	55	43	1.1	Oct. 16, 1965	44	61	3.2
Feb. 21, 1966	42			Feb. 25, 1966	23		
June 3, 1966	22	56	3.0	May 10, 1966	22	58	2.3
Carthage				Elm			
July 31, 1966	34	70	4.5	Aug. 6, 1965	35	83	3.0
Nov. 7, 1966	48			Nov. 6, 1965	33	65	3.7
Feb. 3, 1967	44			Feb. 19, 1966	12		
Apr. 18, 1967	58	65	1.9	May 7, 1966	44	60	1.9
Catron Pond				Enemy Swim			
Aug. 26, 1965	104	52	1.2	July 29, 1965	49	85	5.8
Oct. 18, 1965	44	72	1.3	Nov. 13, 1965	50	57	6.5
Feb. 26, 1966	140			Feb. 11, 1966	16		
May 9, 1966	68	65	3.5	Apr. 25, 1966	11	68	7.1
Cottonwood				Gardner			
Aug. 11, 1965	54	54	0.9	Aug. 26, 1965	53	63	4.9
Oct. 30, 1965	30	63	1.9	Oct. 18, 1965	48	69	3.9
Feb. 20, 1966	80			Feb. 26, 1965	32		
June 4, 1966	42	48	1.4	May 9, 1966	30	71	2.0

Table 2. (continued)

Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at Depth of 1% 4 cm. transmission (in meters)		Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at Depth of 1% 4 cm. transmission (in meters)	
Hayes				Hiddenwood			
Aug. 17, 1966	37	76	3.9	Aug. 29, 1965	65	70	1.6
Oct. 8, 1966	49	60	5.4	Oct. 30, 1965	90	42	0.8
Feb. 13, 1967	56			Feb. 20, 1965	47		
Apr. 12, 1967	51	60	3.0	May 7, 1966	32	60	1.8
Henry				Horse Thief			
Apr. 16, 1966	25	60	3.3	Aug. 8, 1966	23	55	5.1
Nov. 6, 1966	22	80	3.7	Oct. 16, 1966	39		
Feb. 5, 1967	30			Feb. 12, 1967	56		
Apr. 6, 1967	44			Apr. 15, 1967			
Hendricks				Iron Creek			
July 27, 1965	111	71	1.0	Aug. 7, 1966	24	60	5.2
Dec. 2, 1965	55			Oct. 15, 1966	22		
Feb. 19, 1966	61			Feb. 13, 1967	23		
Apr. 22, 1966		50	0.8	Apr. 15, 1967	23	60	2.5
Herman				Isabell			
July 16, 1966	81	52	1.8	Aug. 29, 1965	41	62	1.8
Nov. 4, 1966	82	50	0.9	Nov. 19, 1965	37	62	4.1
Feb. 3, 1967	54			Feb. 27, 1966	16		
Apr. 4, 1967	81			May 8, 1966	22	68	3.5

Table 2. (continued)

Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at      Depth of 1% 4 cm. transmission (in meters)		Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at      Depth of 1% 4 cm. transmission (in meters)	
Kadoka				Mitchell			
Aug. 9, 1966	125		0.3	Aug. 11, 1966	78	57	1.8
Oct. 8, 1966	104			Nov. 7, 1966	38	65	2.0
Feb. 10, 1967	106			Feb. 4, 1967	35		
Apr. 12, 1967	153		0.1	Apr. 7, 1967	49		
Madison				Newell			
July 16, 1966	51	50	0.9	Aug. 25, 1965	12	76	7.0
Nov. 4, 1966	32	60	4.5	Oct. 17, 1965	12	62	4.5
Feb. 3, 1967	100			Feb. 26, 1966			
Apr. 5, 1967	73	50	0.9	May 10, 1966	22	65	5.0
Merindahl				Pactola			
Aug. 15, 1966	67			Aug. 7, 1966	31		
Nov. 6, 1966	29	60	1.9	Oct. 16, 1966	9	87	12.0
Feb. 5, 1967	22			Feb. 12, 1967	31		
Apr. 8, 1967	43			Apr. 15, 1967	21	59	12.1
Mission				Pocasse			
Aug. 10, 1966	50	40	0.5	Aug. 12, 1965	66	29	0.5
Oct. 9, 1966	57	55	2.6	Oct. 31, 1965	115	35	0.5
Feb. 10, 1967	25			Feb. 20, 1966	53		
Apr. 11, 1967	77	40	0.2	June 7, 1966	58	40	0.5

Table 2. (continued)

Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at      Depth of 1% 4 cm. transmission (in meters)		Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at      Depth of 1% 4 cm. transmission (in meters)	
Poinsett				Shadehill			
Aug. 5, 1965	65	69	0.5	Aug. 27, 1965	68	61	3.6
Nov. 7, 1965	51	54	1.9	Oct. 19, 1965	55	66	4.1
Feb. 23, 1966	29			Feb. 27, 1966	16		
June 18, 1966	19	68	4.6	May 8, 1966	28	73	3.2
Punished Woman				Sheridan			
July 14, 1965	47			Aug. 6, 1966	23	64	4.5
Nov. 26, 1965	54			Oct. 16, 1966	24	73	7.0
Feb. 18, 1966	87			Feb. 12, 1967	36		
Apr. 22, 1966	44	72	2.4	Apr. 15, 1967	19	40	4.2
Rose Hill				Stephen Mission			
Aug. 17, 1966		55	1.8	Aug. 17, 1967	62	50	0.5
Oct. 30, 1966	58	65	7.0	Oct. 29, 1966	65		
Feb. 4, 1967	32			Feb. 4, 1967	46		
Apr. 8, 1967	50			Apr. 8, 1967	57		
Roy				Traverse			
July 29, 1965	66	69	3.6	July 13, 1965	81		0.5
Nov. 13, 1965	34			Nov. 11, 1965	74	65	1.4
Feb. 11, 1966	15			Feb. 18, 1966	76		
Apr. 25, 1966	16	67	6.7	Apr. 23, 1966		54	1.8

Table 2. (continued)

Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at    Depth of 1% 4 cm. transmission (in meters)		Lake and Date	Turbidity (J.T.U.)	<u>Light Transmission</u> % at    Depth of 1% 4 cm. transmission (in meters)	
Twin (Sanborn Co.)				White			
Aug. 16, 1966	119		0.5	July 29, 1965	65	72	1.2
Nov. 7, 1966	98	77	1.9	Oct. 12, 1965	27	67	3.7
Feb. 4, 1967	99			Feb. 11, 1966	39		
Apr. 7, 1967	51		1.1	June 3, 1966		59	1.1
Twin (Spink Co.)							
Aug. 7, 1965	87	53	0.5				
Nov. 6, 1965	44	50	1.5				
Feb. 21, 1966	47						
June 4, 1966	68	52	1.8				
Vermillion							
Aug. 15, 1966	62	71	1.9				
Nov. 5, 1966	51	64	1.0				
Feb. 8, 1967	21						
Apr. 5, 1967	43						
Waggoner							
Aug. 9, 1966	39	45	2.6				
Oct. 29, 1966	17	67	4.1				
Feb. 13, 1967	26						
Apr. 12, 1967	42	45	1.4				

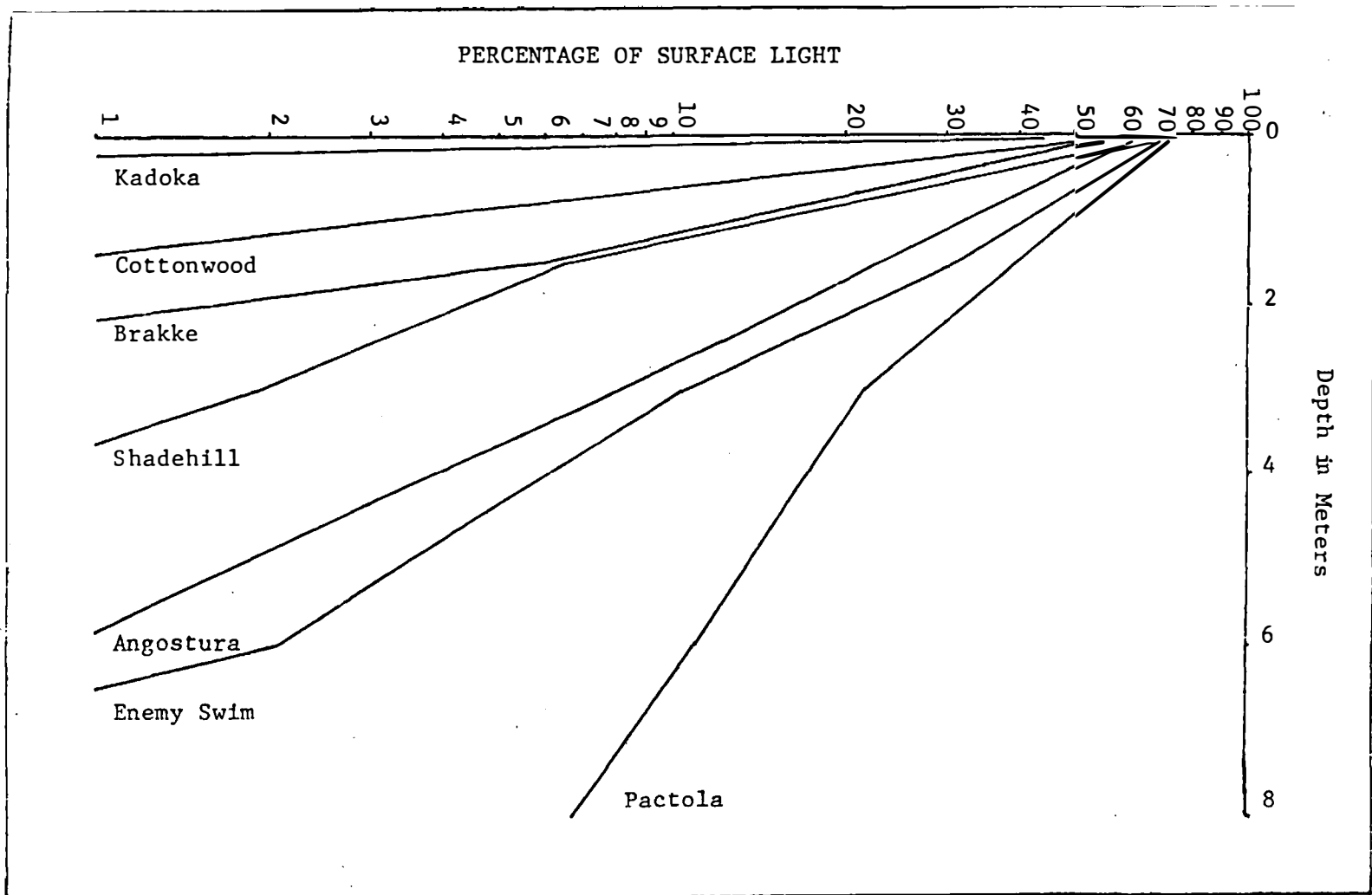


Figure 5. Relation between depth and light transmission, expressed as relative intensity of surface light, for a series of South Dakota lakes.

which 1 per cent of the surface light intensity penetrated varied from 0.2 meters (Kadoka) to 12.0 meters (Pactola). Some of the deeper reservoirs show a slight increase in rate of light transmission with increase in depth (Figure 5). Hutchinson (1957) attributes this phenomenon to highly colored waters which change the spectral composition of the light penetrating the lake. Since these waters did not appear to be highly colored it appears that plankton populations may have reduced the rate of light transmission in the upper few meters of water. A decrease in rate of light penetration with increased depth was evident in a number of shallow lakes. This decrease is probably due to an increase in opacity of bottom waters caused by stirring of the bottom sediments.

## Chemical Considerations

Data obtained from samples taken at all depths were averaged by season for each lake. All data so designated is referred to as an average or seasonal value. This allows seasonal comparisons within each lake. Seasonal values of each lake were averaged to obtain an annual value which permits comparisons among lakes. This value is designated as the mean. Mean values were averaged to obtain representative values for groups of lakes in contrasting physical divisions or geologic areas of the state. These values are referred to as regional means.

### Dissolved Oxygen

The Aquatic Life Advisory Committee (1955) recommended that dissolved oxygen not go below 5 parts per million for more than 8 hours in 24, and never below 3 parts per million unless the population consists of coarse fish which can tolerate dissolved oxygen concentrations as low as 0.5 parts per million. They recommend a minimum of 6 parts per million for coldwater fish habitats. Douderoff and Shumway (1967) accept these standards as sound under present knowledge.

Dissolved oxygen in all lakes was fairly constant from surface to bottom during the spring and fall (Table 3). Lakes Angostura and Hayes showed a pronounced decrease in oxygen with increased depth in the summer only. The following lakes showed a marked decrease in oxygen with increased depth in the winter: Buffalo, Catron Pond, Durkee, Henry, Herman, Hiddenwood, Isabell, Stephen Mission, Traverse,

Table 3. Ranges of dissolved oxygen concentrations from top to bottom of lakes studied, July 1965 to July 1967.

Lake	Summer	Fall	Winter	Spring
Alvin	8.6 - 4.7	10.7 - 10.8	5.0 - 1.4	9.5 - 7.4
Amsden	12.1 - 7.2	11.4 - 11.4	7.2 - 2.7	11.5 - 11.1
Angostura	6.6 - 0.3	8.1 - 8.1	12.0 - 14.0	10.1 - 10.1
Bigstone	11.8 - 7.1	10.5 - 10.4	10.5 - 9.3	14.2 - 13.5
Bismarck	7.2 - 0.0	8.0 - 6.1	4.6 - 0.1	10.1 - 7.0
Brakke	8.2 - 7.8	9.9 - 8.4	13.2 - 10.6	9.6 - 9.6
Buffalo	7.4 - 6.9	12.9 - 12.1	8.5 - 5.1	11.0 - 11.0
Burke	7.9 - 4.4	8.4 - 7.7	1.0 - 0.1	9.7 - 10.0
Byron	6.4 - 5.8	10.4 - 9.8	11.0	7.7 - 7.7
Carthage	5.3 - 3.4	12.3 - 12.1	10.6 - 10.8	9.7 - 9.7
Catron Pond	7.5 - 7.5	8.5 - 8.1	10.5 - 6.5	7.7 - 7.7
Cottonwood	6.1 - 5.8	9.7 - 0.0	8.0 - 7.2	7.5 - 7.0
Durkee	8.0 - 8.0		11.1 - 7.7	9.1 - 9.1
Elm	9.6 - 5.9	11.0 - 11.0	10.8 - 10.6	9.7 - 9.5
Enemy Swim	8.1 - 7.8	10.9 - 10.3	11.5 - 9.5	11.3 - 11.5
Gardner	8.0 - 8.0	9.8 - 9.9	13.4 - 12.0	8.7 - 8.6
Hayes	7.8 - 2.1	9.7 - 7.8	14.0	9.0 - 8.8
Henry	6.0 - 5.1	10.8 - 9.8	11.6 - 4.0	9.7 - 9.5
Hendricks	7.6 - 6.6	15.0 - 14.3	6.9 - 6.8	11.5 - 10.6
Herman	9.6 - 10.4	11.6 - 11.6	4.7 - 0.3	11.0 - 10.7
Hiddenwood		7.6 - 7.0	0.2	9.6 - 5.7
Horse Thief	8.7 - 0.0	6.8	12.5 - 1.2	9.7
Iron Creek	7.5 - 0.0	7.4	3.5 - 0.4	7.4 - 0.0
Isabell	8.3 - 7.9	9.7 - 9.5	12.1 - 5.6	9.4 - 9.4
Kadoka	7.0	9.7	13.5	8.9
Madison	6.1 - 5.6	10.8 - 10.8	16.4 - 18.2	9.7 - 9.3
Merindahl	7.6 - 2.3	12.0 - 11.8	11.5 - 0.9	10.7 - 10.4
Mission	7.5 - 6.0	9.3 - 9.5	10.6 - 11.8	9.9

Table 3. (continued)

Lake	Summer	Fall	Winter	Spring
Mitchell	7.1 - 4.5	11.5 - 11.3	11.3 - 9.5	10.6 - 10.6
Newell	8.1 - 7.5	9.8 - 9.4	13.3 - 10.9	9.2 - 9.5
Pactola	7.2 - 8.0	8.3 - 3.9	10.0 - 3.4	10.4 - 8.4
Pocasse	12.7 - 7.0	9.5 - 9.5	9.3 - 8.2	10.2 - 8.8
Poinsett	8.5 - 8.0	10.2 - 9.8		8.2 - 5.5
Punished Woman	10.9 - 7.9	13.0 - 13.6	13.9 - 13.5	11.9 - 12.2
Rose Hill	8.0 - 0.0	10.8 - 10.8	14.1 - 0.6	8.2 - 7.5
Roy	9.8 - 9.6	10.9 - 10.9	11.6 - 9.1	11.7 - 11.7
Shadehill	7.4 - 7.1	10.0 - 9.8	12.4 - 11.0	10.7 - 11.7
Sheridan		7.5 - 7.5	11.2 - 9.0	10.4 - 8.0
Stephen Mission	8.4 - 7.2	10.8 - 9.7	12.2 - 7.2	9.6 - 9.4
Traverse	10.9 - 10.0	10.7 - 10.3	11.0 - 8.1	12.3 - 12.4
Twin (Sanborn)	9.6 - 10.0	12.3 - 12.0	11.8	10.6
Twin (Spink)	8.3 - 7.8	10.3 - 10.0	10.2 - 7.6	7.1 - 7.1
Vermillion	8.0 - 6.6	12.8 - 11.7	6.0 - 3.2	9.6 - 9.4
Waggoner	10.1 - 9.9	10.8 - 10.6	12.7	9.1 - 8.9
White	11.6 - 11.1	10.9 - 10.9	12.7 - 12.2	8.0 - 7.5

Twin (Spink Co.), and Vermillion. Many lakes developed a pronounced oxygen concentration decrease in both summer and winter. These lakes include: Alvin, Amsden, Bismarck, Burke, Horse Thief, Iron Creek, Merindahl, and Rose Hill. Eighteen of the 45 lakes studied had lower than the recommended oxygen concentrations at some time during the study. Although many lakes did not exhibit marked oxygen stratification or depletion during sampling dates this depletion may have developed during other parts of the seasons.

### pH

Water in all lakes studied was basic, ranging in pH from 7.1 to 11.3 (Table 4). The pH was usually quite uniform from top to bottom or slightly more acid at the bottom. Most lakes which stratified thermally also showed greater variation in pH. Lakes Bismarck, Iron Creek, Rose Hill, and Sheridan exhibited pH differences of over 1.0 from top to bottom.

### Specific Conductance

Specific conductance is a measure of a water's capacity to carry an electric current. This measure may be correlated with salinity or total dissolved solids since salinity is defined as the total concentration of the ionic components.

Rawson (1951) and Northcote and Larkin (1956) attributed much of the difference in plankton, bottom fauna, and fish production between lakes to differences in total dissolved solid content. Most lakes occupying open basins have a total dissolved solid concentration

Table 4. Average pH of lakes studied, July 1965 to July 1967.

Lake	Summer	Fall	Winter	Spring
Alvin	8.5	8.4	7.5	7.9
Amsden	8.5	8.8	8.4	8.7
Angostura	7.7	8.1	8.2	8.1
Bigstone	8.4	8.0	8.5	8.9
Bismarck	7.6	7.1	7.1	7.6
Brakke	8.7	8.7	8.0	8.3
Buffalo	8.6	8.7	8.0	8.4
Burke	9.0	8.4	7.4	8.3
Byron	8.5	8.4	8.7	8.2
Carthage	8.2	8.3	7.9	8.1
Catron Pond	8.8	8.0	8.4	8.2
Cottonwood	9.0	8.7	8.8	8.7
Durkee	8.9	8.2	8.5	8.0
Elm	8.8	8.7	8.7	7.8
Enemy Swim	8.4	8.4	8.4	8.4
Gardner	8.3	8.9	8.7	8.1
Hayes	8.2	8.3	8.4	8.1
Henry	8.3	8.1	7.4	8.3
Hendricks	8.5	8.7	8.3	8.4
Herman		8.7	7.9	8.7
Hiddenwood	8.7	7.7	8.3	7.9
Horse Thief	7.4	6.8	7.2	7.2
Iron Creek	7.7	7.6	7.4	7.3
Isabell	8.0	8.5	8.8	7.9
Kadoka	8.6	8.6	8.3	8.5
Madison		11.3	8.4	8.3
Merindahl	8.0	8.5	7.6	8.2
Mission	8.4	8.6	7.8	8.6

Table 4. (continued)

Lake	Summer	Fall	Winter	Spring
Mitchell	8.2	8.3	8.0	8.3
Newell	8.4	8.1	8.3	7.8
Pactola	7.9	8.0	8.3	8.3
Pocasse	9.1	7.8	8.5	7.8
Poinsett	8.9	8.5	8.6	8.7
Punished Woman	8.8	8.3	8.2	8.0
Rose Hill	8.2	8.7	8.0	7.6
Roy	8.7	9.0	8.7	8.7
Shadehill	8.4	8.6	8.8	7.9
Sheridan	7.8	7.4	8.4	8.2
Stephen Mission	8.6	8.4	7.9	8.0
Traverse	8.2	8.0	8.2	8.6
Twin (Sanborn)	8.6	7.8	7.7	8.2
Twin (Spink)	9.0	8.9	8.9	8.7
Vermillion	9.3	9.0	7.8	8.1
Waggoner	9.6	8.7	8.3	8.1
White	8.5	7.8	8.0	

of between 100 and 200 ppm (Reid, 1961). Evaporation from lakes with closed basins may concentrate dissolved solids to over 100,000 ppm causing pronounced changes in species and numbers of fauna and flora present.

The American Public Health Association (1965) stated: "The amount of dissolved ionic matter in a sample may often be estimated by multiplying the specific conductance by an empirical factor. This factor may vary from 0.55 to 0.9, depending on the soluble components of the particular water and the temperature of the measurement. Relatively high factors may be required for saline or boiler waters, whereas lower factors may apply where considerable hydroxide or free acid is present."

The correction of specific conductance measurements to specific conductance at a standard temperature is accomplished by applying a temperature coefficient. Smith (1962) gave this coefficient as a two per cent rate of change of conductance for each degree centigrade temperature change in dilute aqueous solutions.

Specific conductance in shallow lakes (maximum depth less than 15 feet) increased from a spring or summer minimum to a winter maximum (Table 5). This winter maximum was more pronounced in lakes of northern South Dakota (e.g. Bigstone, Buffalo, Gardner, and Hiddenwood). The deeper lakes in this area (e.g. Amsden, Pocasse, Roy, and White) also showed maximum specific conductance during the winter while deeper lakes in the southern part of the state did not. This maximum conductance was probably due to concentration of mineral

Table 5. Average specific conductance of lake waters (micromhos per centimeter at 25 C), July 1965 to July 1967.

Lake	Summer	Fall	Winter	Spring
Alvin	503	442	440	363
Amsden	865	1288	1520	
Angostura	1700	2200	1620	1700
Bigstone	655	853	1093	
Bismarck	103	118	118	112
Brakke	493	309	450	383
Buffalo	580	505	887	568
Burke	290	245	290	263
Byron	2400	3900	4800	1750
Carthage	340	305	362	320
Catron Pond	408	530	790	185
Cottonwood	2867	2600	4750	1700
Durkee	196	265	359	225
Elm	599	747	661	230
Enemy Swim	420	439	571	465
Gardner	517	767	973	120
Hayes	333	273	320	301
Henry	1307	957	940	1000
Hendricks	527	850	650	
Herman	750	908	840	685
Hiddenwood	453	593	970	180
Horse Thief	70	68	53	82
Iron Creek	106	140	112	113
Isabell	318	433	513	80
Kadoka	368	380	450	400
Madison	1280	1175	1375	1055
Merindahl	827	875	840	1650
Mission	268	194	236	180

Table 5. (continued)

Lake	Summer	Fall	Winter	Spring
Mitchell	527	870	790	785
Newell	508	468	880	470
Pactola		247	221	226
Pocasse	510	476	883	165
Poinsett	780	1282	1233	840
Punished Woman	430	383	700	380
Rose Hill	313	234	268	232
Roy	1150	1210	1518	1165
Shadehill	1055	1700	1730	543
Sheridan	173	157	160	172
Stephen Mission	367	423	475	393
Traverse	905	1025	1650	
Twin (Sanborn)	3650	3700	3200	2800
Twin (Spink)	3225	4400	6000	2117
Vermillion	420	522	555	335
Waggoner	440	447	655	535
White	800	997	1175	

elements under the ice, and was more pronounced in the northern area where ice depth was greatest. The pronounced decrease in specific conductance in north central and northwestern South Dakota during spring 1966 may be attributed to increased spring runoff caused by a severe March blizzard. The U. S. Department of Commerce (1966) showed that precipitation over much of this area deviated +2.0 inches or more from normal during the first three months of 1966. Resulting increased runoff could cause dilution or flushing in lakes which normally receive limited precipitation. Lakes affected were: Byron, Catron Pond, Cottonwood, Elm, Gardner, Hiddenwood, Isabell, Pocasse, Shadehill, and Twin (Spink Co.).

The mean specific conductance of lakes studied, and of lakes studied by the U. S. Department of Interior (1964 and 1965) are plotted in Figure 6. Variation in specific conductance among lake waters was usually associated with physical divisions of the state and substages of the Wisconsin glacial age. Most lakes east of the Missouri River which lie in soil deposited by the Mankato substage had much higher mean specific conductance (330 to 1260 micromhos at 25 C) than did lakes in the Cary and Iowan substages (260 to 880 micromhos at 25 C). Exceptions were lakes occupying closed basins or overlying aquifers (Figure 6). These lakes had increased mean specific conductance ranging from 1780 to 3940 micromhos at 25 C in the Mankato, and from 800 to 16,000 micromhos at 25 C in the Cary and Iowan substages.

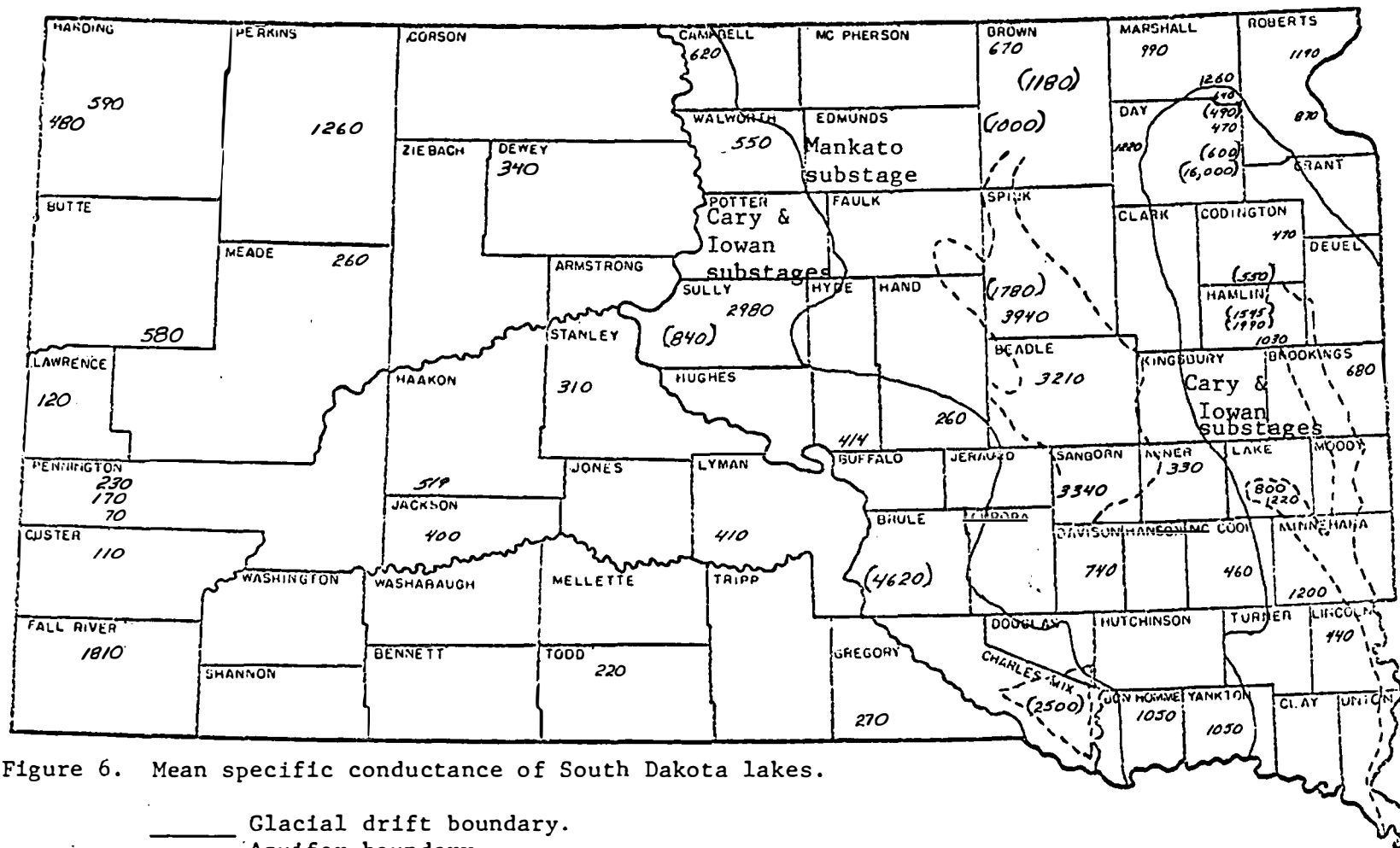


Figure 6. Mean specific conductance of South Dakota lakes.

— Glacial drift boundary.

- - - Aquifer boundary.

Values in parenthesis indicate U. S. Department of Interior data.

Lakes in the unglaciated portion of the state, west of the Missouri River, generally had lower mean specific conductance values than did lakes east of the river. The lowest mean specific conductance readings were obtained in the Black Hills (70 to 230 micromhos at 25 C). The next lowest values were in lakes of the Southern Plateaus ranging from 220 to 270 micromhos at 25 C. The remaining lakes in the unglaciated area, except for Shadehill and Angostura, ranged from 310 to 590 micromhos at 25 C. The high mean specific conductance of Angostura, 1810 micromhos at 25 C, is attributed to the influent Cheyenne River. Water from the Cheyenne River has a higher dissolved solids content than the water from any of the other major streams in South Dakota (U. S. Geological Survey and U. S. Bureau of Reclamation, 1964). Drainage from irrigation projects contributes much of the dissolved solids in this water.

### Alkalinity

Ball (1948), Carlander (1955), and Turner (1960) found that lakes with high alkalinities usually have higher standing crops of fish than do lakes with low alkalinities. Moyle (1946) states that where alkalinity is above 40 ppm there seems to be no relationship between increased carbonate and yield. He considers this a natural separation point between soft and hard water.

Alkalinity is usually imparted by hydroxide, carbonate, and bicarbonate. Hydroxide alkalinity was not found in any of the lakes studied but carbonate alkalinity occurred commonly. Alkalinity

relationships were calculated stoichiometrically (carbonate =  $2X$  phenolphthalein and bicarbonate = total - carbonate) from the phenolphthalein and bromo cresol green-methyl orange alkalinities. Carbonate and bicarbonate alkalinities, as mg/l  $\text{CaCO}_3$ , were averaged seasonally and presented in Table 6.

Little variation was found in the seasonal pattern of alkalinity values among lakes. They generally increased from a spring minimum to a winter maximum. Variations in total alkalinity from top to bottom of individual lakes were usually small. In deep lakes where there is a thermocline total alkalinity is usually greater below than above the thermocline (Moyle, 1956). Several lakes showed a marked increase in total alkalinity from top to bottom during the summer. The increase in the alkalinity of bottom waters of these lakes is indicated below: Angostura, 17 mg/l; Bismarck, 19 mg/l; Horse Thief, 17 mg/l; Iron Creek, 58 mg/l; Rose Hill, 83 mg/l; and Sheridan, 21 mg/l.

Mean total alkalinity west of the Missouri River ranged from 38-264 mg/l. Alkalinity was lowest in the Black Hills, ranging from 38 to 141 mg/l. Horse Thief was the only soft water lake with 38 mg/l  $\text{CaCO}_3$ . Lakes of the Southern Plateaus had a small mean total alkalinity range of 188-210 mg/l. Lakes in the Pierre Hills area ranged from 85 mg/l in the west to 177 mg/l in the east. Lakes of the Northern Plateaus ranged from 123 to 214 mg/l with highest values in the northwest which has a high evaporation to rainfall ratio.

Table 6. Average carbonate and bicarbonate alkalinities (as mg/l  $\text{CaCO}_3$ ) of lakes studied, July 1965 to July 1967.

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>	
	Car- bon- ate	Bi- car- bonate	Car- bon- ate	Bi- car- bonate	Car- bon- ate	Bi- car- bonate	Car- bon- ate	Bi- car- bonate
Alvin	13	101	0	113	0	157	0	131
Amsden	48	130	30	158	26	204	22	169
Angostura	1	112	1	107	3	116	0	113
Bigstone	47	100	0	171	12	128		
Bismarck	1	59	0	55	0	165	0	54
Brakke	20	129	17	169	0	192	5	163
Buffalo	36	175	21	211	0	291	7	182
Burke	55	126	16	210	0	231	7	191
Byron	7	129	20	303	64	304	22	206
Carthage	4	135	5	166	0	200	0	137
Catron Pond	113	117	0	216	11	265	10	197
Cottonwood	192	340	110	437	122	848	60	315
Durkee	44	99	3	105	8	134	0	96
Elm	37	134	9	180	23	103	0	117
Enemy Swim	45	194	10	197	0	260	6	207
Gardner	50	190	33	184	29	278	6	209
Hayes	6	95	5	124	8	169	0	143
Henry	14	170	0	165	0	221	6	194
Hendricks	17	134	20	186	8	112	9	136
Herman	91	77	26	197	0	300	43	176
Hiddenwood	60	120	0	165	14	201	7	132
Horse Thief	0	32	0	39	0	44	0	35
Iron Creek	7	84	0	73	0	94	0	91
Isabell	0	146	1	129	11	167	0	89
Kadoka	10	166	16	119	10	275	20	203

Table 6. (continued)

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>	
	Car- bon- ate	Bi- car- bonate	Car- bon- ate	Bi- car- bonate	Car- bon- ate	Bi- car- bonate	Car- bon- ate	Bi- car- bonate
Madison	39	117	13	182	10	237	4	171
Merindahl	1	147	12	114	0	156	4	139
Mission	13	126	13	178	0	243	9	167
Mitchell	9	131	6	129	0	167	4	153
Newell	13	74	0	81	6	75	0	88
Pactola	0	131	0	139	0	148	0	147
Pocasse	67	104	0	195	22	235	0	92
Poinsett	75	154	9	213	27	225	67	135
Punished Woman	39	117	8	175	15	140	10	140
Rose Hill	7	124	10	100	3	149	0	100
Roy	54	209	44	182	42	262	24	218
Shadehill	32	268	19	251	21	226	10	225
Sheridan	7	88	0	100	2	99	0	99
Stephen Mission	21	93	3	125	0	162	0	112
Traverse	43	131	0	191	11	171	20	150
Twin (Sanborn)	23	66	0	94	0	127	2	109
Twin (Spink)	194	170	118	320	149	466	80	259
Vermillion	60	68	29	92	0	137	0	115
Waggoner	51	21	0	79	10	100	0	87
White	34	133	0	168	0	216	4	178

Patterns of total alkalinity concentrations were not as evident east of the Missouri River. Alkalinity of lakes with open basins was generally highest in the north (160-260 mg/l) and lowest in the south (100-160 mg/l) where rainfall was greater and carbonate leaching deeper. This relationship is modified in those lakes occupying closed basins or overlying aquifers as they have high total alkalinities (194-606 mg/l) as well as high specific conductance.

### Sulfate

Sulfate is important in natural waters in many ways. Sulfur is an essential component of protoplasm and is necessary for plant growth. Lack of sulfate can inhibit development of phytoplankton populations thus limiting production (Reid, 1961). Moyle (1945) pointed out that there is a noticeable biological relationship between concentration of sulfates and distribution of aquatic plants in Minnesota. Rawson and Moore (1944) studied lakes in Saskatchewan with high sulfate salinity and found a decrease in number of fish species when salinity exceeded 7,000 ppm.

Eddy (1963) stated that waters in southwestern and extreme western Minnesota and extending westward into the Dakotas are extremely high in dissolved salts. In these waters the concentration of sulfate salts (as sulfate ions) often exceeds the concentration of carbonates (as carbonate ion). Moyle (1945) found that in Minnesota such waters typically have a sulfate ion concentration greater than

150 ppm. He states that waters high in sulfates are usually high in carbonates, nitrates, and phosphates.

Seasonal concentrations of sulfate ions for South Dakota lakes are shown in Table 7. Sulfate ion concentrations did not show any consistent seasonal patterns among lakes. Some lakes had maximum concentrations of sulfates in the winter while others reached their maximum value in the spring, summer, or fall. Areas where concentrations of sulfate salts (as sulfate ions) exceed carbonate salts (as carbonate ions) in lakes are shown in Figure 7. Carbonate exceeded sulfate in the region west of the Missouri River except for portions of the Pierre Hills drained by the main tributaries of the Belle Fourche and Cheyenne Rivers and part of the Grand River drainage. In lakes east of the Missouri River concentrations of sulfate exceeded carbonate except in a small area of the Coteau des Prairie.

Lakes west of the Missouri River were lowest in sulfate ion concentration. The lakes in the Black Hills and the Southern Plateaus divisions had a regional mean concentration of 23 mg/l sulfate ion. Angostura reservoir had a considerably higher mean concentration of 689 mg/l. Patterns were not evident in the Pierre Hills or Northern Plateaus. The regional mean sulfate ion concentration of lakes in these divisions was 116 mg/l.

Lakes east of the Missouri River had consistently higher sulfate ion concentration than those west of the river. Highest concentrations were found in lakes lying in Mankato drift (Figure 6). Lakes

Table 7. Average sulfate ion concentrations (mg/l) of lakes studied, July 1965 to July 1967.

Lake	Summer	Fall	Winter	Spring	Mean
Alvin	145	130	160	105	135
Amsden	450	176	175		207
Angostura	701	765	565	725	689
Bigstone	225	136	310		224
Bismarck	10	11	12	13	12
Brakke	74	72	112	87	86
Buffalo	225	92	250	115	171
Burke	12	5	37	15	17
Byron	1000	400	1950	210	890
Carthage	67	78	45	60	63
Catron Pond	65	87	30	35	54
Cottonwood	1030	600	855	425	728
Durkee	25	18	45	90	45
Elm	152	112	100		121
Enemy Swim	125	91	49	55	80
Gardner	136	210	80		142
Hayes	75	65	82	70	73
Henry	387	287	347	385	352
Hendricks	325	142	186		218
Herman	227	280	240	238	246
Hiddenwood	80	125	140		115
Horse Thief	10	8	12	20	13
Iron Creek	3	8	27	0	10
Isabell	40	106	50	10	52
Kadoka	35	45	68	55	51
Madison	400	282	537	392	403
Merindahl	345	352	450	400	387
Mission	20	5	10	7	11

Table 7. (continued)

Lake	Summer	Fall	Winter	Spring	Mean
Mitchell	248	352	337	287	306
Newell	150		230		190
Pactola	50	48	43	49	48
Pocasse	160	150	155		155
Poinsett	290	116	275	207	222
Punished Woman	150	60	120	120	113
Rose Hill	32	47	54	38	43
Roy	450	600	640	390	520
Shadehill	348	160	300	65	218
Sheridan	18	25	25	32	25
Stephen Mission	123	140	140	122	131
Traverse	250	240	448		313
Twin (Sanborn)	2100	2000	1350	1850	1825
Twin (Spink)	850	491	285	323	487
Vermillion	140	145	210	105	150
Waggoner	100	182	245	192	180
White	325	150	420		298

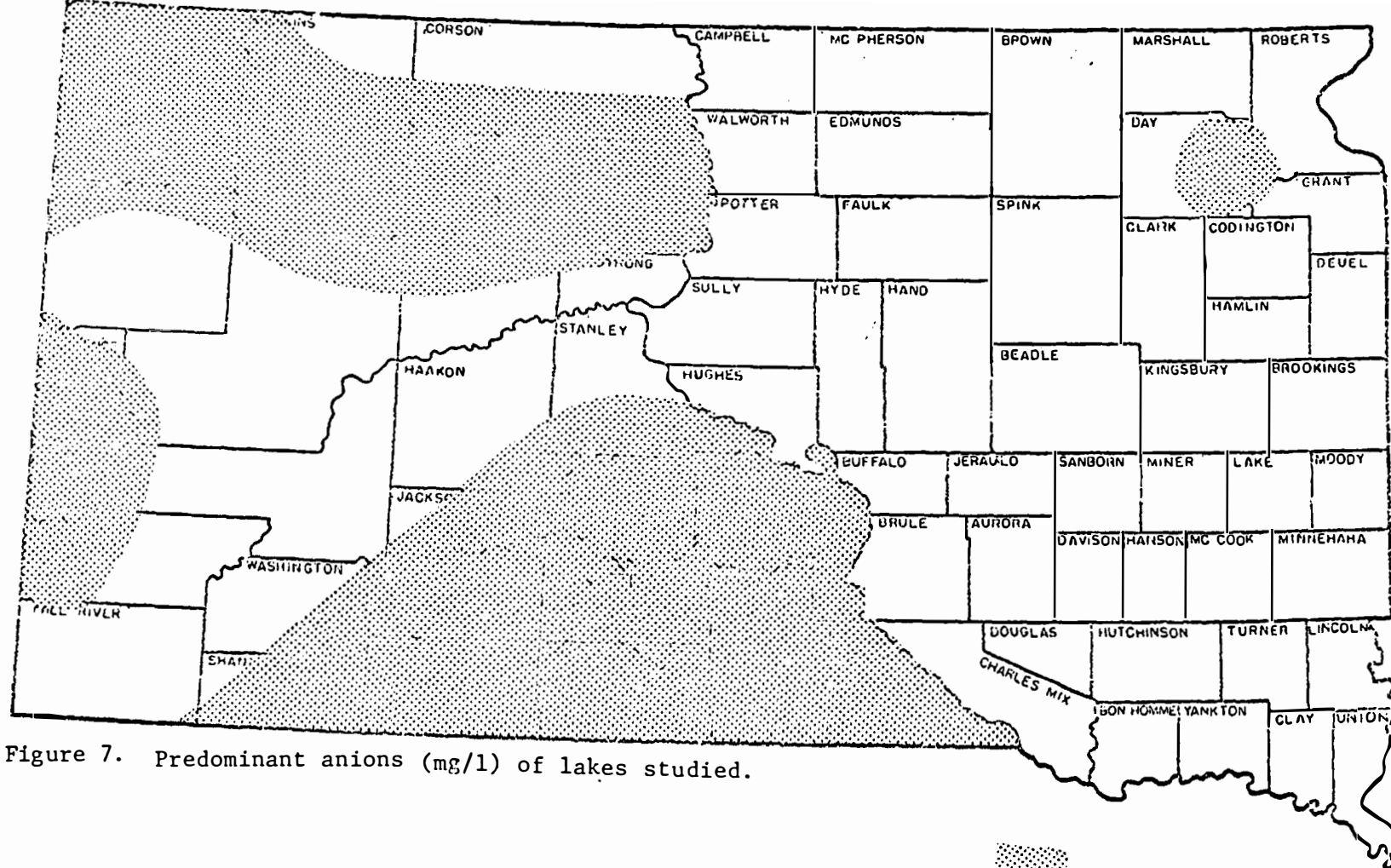


Figure 7. Predominant anions (mg/l) of lakes studied.

occupying closed basins or overlying aquifers in this area had higher mean concentrations ranging from 487 to 1825 mg/l. Lakes with open basins in the Mankato drift had a regional mean concentration of 273 mg/l.

Lakes in the Cary and Iowan till generally had lower concentrations of sulfate ions. Mean concentration of those occupying closed basins or overlying aquifers ranged from 135 to 728 mg/l while the other lakes in the Cary and Iowan till had a regional mean of 128 mg/l sulfate ion.

### Chloride

Moyle (1956) stated that high concentrations of sulfates and chlorides usually occur together. It has long been known that high concentrations of chlorides are related to plant distribution. This may be an osmotic effect since Gerloff, Fitzgerald, and Skoog (1950) demonstrated that there are optimum osmotic concentrations for plants.

High concentrations of chlorides in areas where chloride concentrations are normally low may indicate pollution. Purified effluent liquids from sewage treatment plants tend to be high in chloride (Moyle, 1956). One lake with unusually high chloride content is Lake Madison in Lake County. This lake receives sewage effluent from the treatment plant at the city of Madison. Mean annual chloride content in this lake was 113 mg/l. Lake Herman which is in the same drainfield had a mean chloride concentration of 5 mg/l.

Mean chloride concentration for this series of analyses, ranged from 0.6 to 518.6 mg/l. Thirty-three of forty-five lakes examined had less than 15.0 mg/l chloride (Table 8). Lakes west of the Missouri River were lowest in chloride ion concentration. Lakes in the Black Hills had a mean regional concentration of 1.9 mg/l. The remaining lakes west of the Missouri River, except Angostura, had a regional mean concentration of 4.9 mg/l. Angostura reservoir had a mean concentration of 129.4 mg/l.

Lakes east of the Missouri River had consistently higher chloride ion concentrations than did lakes west of the river. Highest concentrations were found in the Mankato drift of the Wisconsin ice age (Figure 6). Lakes overlying aquifers (Figure 6) or occupying closed basins had higher concentrations than lakes occupying open basins. Among these are: Byron, 177.3 mg/l; Cottonwood, 136.8 mg/l; Twin (Sanborn Co.), 149.0 mg/l; and Twin (Spink Co.), 518 mg/l. The U. S. Geologic Survey and U. S. Bureau of Reclamation (1964) gave results from chemical analyses of waters from principle aquifers in the state. According to these analyses glacial drift aquifers probably have a greater range in chemical quality than any of the other principle aquifers.

Lakes occupying open basins in the Mankato drift had a regional mean concentration of 21.4 mg/l chloride ion, while similar lakes in soils deposited by the Cary and Iowan substages had a regional mean concentration of 5.3 mg/l.

Table 8. Mean chloride ion concentrations (mg/l) of lakes studied, July 1965 to July 1967.

Lake	Mean	Lake	Mean
Alvin	4.9	Isabell	3.3
Amsden	14.0	Kadoka	9.2
Angostura	129.4	Madison	113.0
Bigstone	13.8	Merindahl	5.1
Bismarck	2.5	Mission	1.9
Brakke	11.7	Mitchell	22.6
Buffalo	5.6	Newell	5.5
Burke	5.4	Pactola	1.5
Byron	177.3	Pocasse	5.3
Carthage	4.9	Poinsett	41.5
Catron Pond	1.0	Punished Woman	6.9
Cottonwood	136.8	Rose Hill	3.8
Durkee	3.9	Roy	12.8
Elm	23.9	Shadehill	5.8
Enemy Swim	5.6	Sheridan	3.5
Gardner	0.6	Stephen Mission	9.8
Hayes	4.6	Traverse	15.7
Henry	70.5	Twin (Sanborn)	149.0
Hendricks	4.2	Twin (Spink)	518.6
Herman	5.0	Vermillion	3.5
Hiddenwood	4.5	Waggoner	8.3
Horse Thief	1.2	White	14.5
Iron Creek	0.6		

## Hardness Relationships

Total hardness measures the divalent cations capable of reacting with soap to form precipitates. The principle divalent cations are calcium and magnesium. Temporary hardness or alkalinity are indicators of the carbonate and bicarbonate content of waters expressed as calcium carbonate. Since alkalinity is an indicator of productivity in aquatic ecosystems it has been discussed separately. Non-carbonate hardness (total hardness-alkalinity) is considered to be primarily due to sulfates, chlorides, and nitrates of the metallic ions causing hardness. Negative non-carbonate hardness indicates that more bicarbonate ions are present than are needed to satisfy the divalent metallic ions present. These excess bicarbonate ions are related to monovalent metallic ions, such as sodium and potassium.

Seasonal and mean concentrations of total and non-carbonate hardness, as mg/l calcium carbonate, are presented in Table 9. Both total and non-carbonate hardness tended to be at a maximum during the winter.

Those portions of the state with a predominance of carbonate and bicarbonate as the major anions (Figure 7) usually had these ions in excess of the calcium and magnesium equivalents. Negative non-carbonate hardness was associated primarily with sodium and potassium. Areas with positive non-carbonate hardness paralleled those portions of the state where sulfate and chloride were predominant anions (Figure 7).

Table 9. Average concentrations of total and non-carbonate hardness (as mg/l CaCO<sub>3</sub>) from lakes studied, July 1965 to July 1967.

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	To- tal	Non- car- bonate	To- tal	Non- car- bonate	To- tal	Non- car- bonate	To- tal	Non- car- bonate	To- tal	Non- car- bonate
Alvin	255	141	178	65	229	72	196	65	215	86
Amsden	250	72	274	86	573	343			366	167
Angostura	617	504	679	571	717	598	753	640	692	578
Bigstone	196	49	268	97	371	231			278	126
Bismarck	46	-14	48	-7	51	-114	51	-3	49	-34
Brakke	117	-32	83	103	88	-104	133	-35	105	-68
Buffalo	198	-13	210	-22	446	155	279	90	283	52
Burke	115	-66	130	-96	131	-100	126	-72	126	-83
Byron	370	234	474	151	980	612	364	136	547	283
Carthage	162	23	144	-27	181	-19	153	16	160	-2
Catron Pond	32	198	44	-172	100	-176	48	-159	56	-176
Cottonwood	384	-148	410	-137	644	-326	290	-85	432	-174
Durkee	26	-117	57	-51	79	-63	32	-64	49	-74
Elm	106	-65	187	-2	178	52	76	-41	137	-14
Enemy Swim	140	-99	182	-25	291	31	233	20	213	-18
Gardner	36	-204	42	-175	59	-248	26	-189	41	-201
Hayes	93	-8	85	-44	124	-53	103	-40	101	-36
Henry	618	434	327	162	363	142	434	234	436	243
Hendricks	186	35	203	-3	527	407			305	147
Herman	331	163	485	262	690	390	409	190	479	251
Hiddenwood	94	-86	131	-34	198	-17	60	-79	121	-54
Horse Thief	25	-7	32	-7	34	-10	32	-3	31	-7
Iron Creek	48	-43	62	-11	54	-40	58	-33	56	-32
Isabell	30	-116	87	-43	82	-96	24	-65	56	-80

Table 9. (continued)

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	To- tal	Non- car- bonate	To- tal	Non- car- bonate	To- tal	Non- car- bonate	To- tal	Non- car- bonate	To- tal	Non- car- bonate
Kadoka	45	-131	50	-85	56	-229	52	-171	51	-154
Madison	503	347	513	318	561	314	875	700	613	420
Merindahl	361	213	336	210	368	212	415	272	370	227
Mission	89	-50	91	-100	101	-142	80	-96	90	-97
Mitchell	259	119	290	155	359	192	433	276	335	186
Newell	98	11	139	58	187	106	145	57	142	58
Pactola	109	-22	119	-20	113	-35	129	-18	118	-24
Pocasse	110	-61	115	-80	223	-34	54	-38	126	-53
Poinsett	213	-16	277	55	477	225	319	117	322	95
Punished Woman	144	-12	158	-25	267	113	144	-6	178	18
Rose Hill	124	-7	94	-16	128	-24	101	1	112	-11
Roy	422	159	433	207	768	464	599	357	556	297
Shadehill	74	-226	76	-194	134	-113	59	-176	86	-177
Sheridan	69	-26	89	-11	72	-29	83	-16	79	-20
Stephen Mission	169	55	196	68	265	103	217	105	212	83
Traverse	288	114	307	116	525	343			373	192
Twin (Sanborn)	800	711	812	718	770	643	1252	1141	909	803
Twin (Spink)	435	71	805	367	1072	457	443	104	689	250
Vermillion	226	98	193	72	328	191	236	121	246	121
Waggoner	118	46	168	89	228	118	244	157	190	102
White	260	93	304	136	515	299			360	177

## Calcium and Magnesium

These two elements usually constitute the most abundant ions in fresh waters. Calcium is usually the most abundant of these two ions. Reid (1961) stated that in soft waters calcium makes up, on the average, about 48 per cent and magnesium about 14 per cent of the total cations present. In average hard waters the proportion of magnesium to calcium increases, giving approximately 53 per cent of calcium and 34 per cent of magnesium.

Moyle (1946) indicated that calcium is not a critical factor in the productivity of waters having a total alkalinity greater than 15 ppm. Gerloff et al. (1950) found that calcium concentrations of 9.8 ppm and magnesium concentrations of 0.13 ppm allowed for optimum growth of Coccochloris peniocystris.

Seasonal and mean annual concentrations of calcium and magnesium are presented in Table 10. These ions were usually more concentrated under ice cover during the winter months particularly in lakes of northern South Dakota. Figure 8 shows the areas of South Dakota where calcium and magnesium ions in lakes are predominant over sodium and potassium ions. This predominance was based on weight/volume measurements.

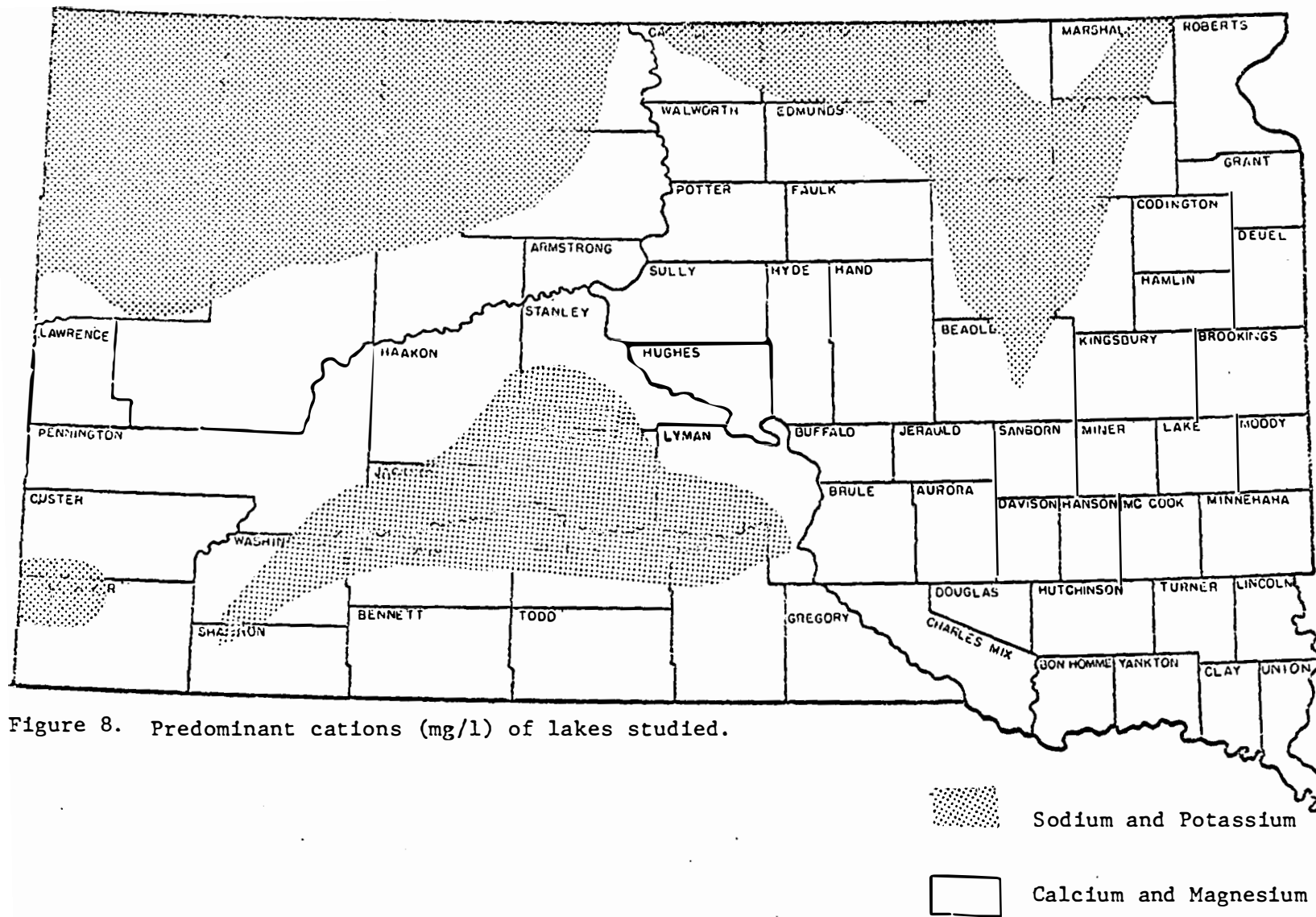
Lakes west of the Missouri River usually had lower concentrations of calcium and magnesium than did lakes east of the river. Mean concentrations of calcium for lakes of the Black Hills, Northern Plateaus, and Southern Plateaus was usually less than 20 mg/l with magnesium concentrations less than 10 mg/l. Other lakes west of the Missouri

Table 10. Average concentrations of calcium and magnesium (mg/l) from lakes studied, July 1965 to July 1967.

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium
Alvin	43.2	35.7	21.6	30.1	46.5	27.3	40.3	23.2	37.9	29.1
Amsden	29.6	42.8	42.9	36.5	48.3	55.9	45.4	119.7	41.6	63.7
Angostura	172.2	45.4	181.3	55.1	204.1	50.3	203.4	59.6	190.3	52.6
Bigstone	22.4	34.0	40.3	40.7	44.8	63.0			35.8	45.9
Bismarck	12.4	3.7	11.2	4.9	12.1	5.0	13.6	4.1	12.3	4.4
Brakke	21.7	15.3	16.0	10.5	17.2	10.9	26.0	16.6	20.2	13.3
Buffalo	20.0	36.0	20.4	38.7	56.9	73.8	39.7	43.7	34.3	48.1
Burke	34.0	8.2	35.2	10.2	38.0	8.8	35.1	9.3	35.6	9.1
Byron	27.2	73.4	35.6	93.6	88.1	184.7	28.4	71.2	44.8	105.7
Carthage	38.8	15.9	27.4	18.3	31.2	25.1	37.7	14.3	33.8	18.4
Catron Pond	8.0	2.9	6.4	6.3	20.4	11.9	12.0	4.4	11.7	6.4
Cottonwood	47.2	64.6	38.4	76.3	37.2	133.9	30.0	52.5	38.2	81.8
Durkee	10.4	2.9	22.7	9.7	31.4	9.0	22.4	5.8	21.7	6.9
Elm	21.6	12.6	22.7	31.6	32.1	20.5	19.2	6.8	23.9	17.9
Enemy Swim	14.4	25.3	18.8	32.7	38.8	47.0	34.6	35.5	26.7	35.1
Gardner	4.8	5.8	5.9	6.6	11.5	7.3	8.8	1.0	7.8	5.2
Hayes	19.1	11.1	15.1	11.4	22.4	16.8	27.1	8.7	20.9	12.0
Henry	57.9	115.0	56.6	45.0	59.3	52.3	66.0	65.4	60.0	69.4
Hendricks	32.0	25.8	41.6	24.1	42.4	102.3			38.7	50.7
Herman	53.3	48.2	59.3	81.9	62.1	130.0	48.4	69.4	55.8	82.4
Hiddenwood	20.0	10.7	22.4	18.2	24.0	33.5	14.4	5.8	20.2	17.1
Horse Thief	6.4	2.3	6.4	3.9	5.3	4.8	8.8	2.4	6.7	3.4
Iron Creek	10.1	5.4	12.0	7.8	11.7	5.9	12.0	6.8	11.5	6.5
Isabell	8.8	1.9	10.7	14.6	20.0	7.8	7.2	1.5	11.7	6.5

Table 10. (continued)

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium	Cal- cium	Mag- ne- sium
Kadoka	10.4	4.7	13.6	3.9	8.0	8.7	12.0	5.3	11.0	5.7
Madison	68.4	80.7	62.5	86.8	81.7	86.7	98.1	153.1	77.7	101.9
Merindah1	109.2	21.5	88.4	28.0	113.2	24.1	112.9	30.1	105.9	25.9
Mission	18.9	10.2	18.4	10.9	22.4	10.9	13.6	11.2	18.3	10.8
Mitchell	48.0	33.7	52.9	38.4	74.5	42.2	82.5	55.2	64.5	42.4
Newell	28.0	6.8	39.2	10.0	45.6	17.6	39.9	11.5	37.9	11.5
Pactola	14.9	17.4	16.8	16.3	15.5	18.1	19.7	19.3	16.7	17.8
Pocasse	38.4	3.4	22.2	14.4	41.6	28.8	14.8	4.1	29.3	12.7
Poinsett	17.6	46.2	23.7	52.8	38.4	92.5	32.8	57.5	28.1	62.3
Punished Woman	16.8	24.8	21.6	25.3	47.7	35.9	26.4	19.0	28.1	26.3
Rose Hill	20.3	17.9	15.5	12.9	20.1	18.9	27.2	7.9	21.0	14.4
Roy	22.4	88.9	24.5	90.2	49.0	156.9	33.6	125.2	32.4	115.3
Shadehill	11.2	11.2	10.4	12.2	22.3	17.7	11.7	7.1	13.9	12.1
Sheridan	14.9	7.7	17.6	11.0	13.5	9.6	18.4	8.9	16.1	9.3
Stephen Mission	39.5	16.8	42.8	21.6	50.9	33.5	51.9	21.2	46.3	23.3
Traverse	28.0	53.0	52.1	43.0	74.1	82.7			51.4	59.6
Twin (Sanborn)	320.3	264.3	324.9	198.9	308.3	298.9	501.3	201.2	363.7	240.8
Twin (Spink)	19.2	94.0	15.1	186.4	24.8	245.5	20.8	94.9	20.0	155.2
Vermillion	40.0	30.6	26.4	30.8	49.2	49.9	37.4	34.7	38.3	36.5
Waggoner	24.0	13.9	42.0	15.1	76.1	23.5	67.7	17.9	52.5	17.6
White	40.0	33.9	52.9	41.8	81.9	77.3			58.3	51.0



River ranged from about 20 to 50 mg/l calcium, while magnesium concentrations in this area ranged from 10 to 20 mg/l. Highest concentrations of calcium and magnesium west of the Missouri were in Angostura reservoir with 190 and 52 mg/l respectively.

Lakes east of the Missouri River generally had higher concentrations of these two elements than those west of the river. Highest concentrations were found in lakes lying in the Mankato drift, with mean calcium and magnesium concentrations ranging from 20-364 mg/l and 18-241 mg/l respectively. Lakes in the Cary and Iowan till had lower ranges of calcium and magnesium varying from 20-78 and 13-102 mg/l respectively.

#### Sodium and Potassium

Sodium and potassium concentrations usually follow calcium and magnesium in waters of open river systems. Reid (1961) stated that where sodium and potassium occur in low concentrations the proportion of sodium is only slightly greater than that of potassium. As the total content of both increases the concentration of sodium greatly exceeds that of potassium.

Concentrations of sodium and potassium varied widely throughout the state. Mean concentrations of sodium ranged from 3-558 mg/l, while potassium ranged from 1-51 mg/l. Seasonal and mean annual concentrations of these elements are presented in Table 11. The areas in which sodium and potassium are the predominant cations are shown in Figure 8.

Table 11. Average concentrations of sodium and potassium (mg/l) from lakes studied, July 1965 to July 1967.

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium
Alvin	15.0	12.3	15.5	12.1	17.0	12.6	13.0	11.7	15.1	12.2
Amsden	91.2	12.6	18.6	12.1	115.2	11.0			75.0	11.9
Angostura	252.0	8.2	240.0	8.7	220.0	11.0	130.0	7.9	210.5	9.0
Bigstone	40.8	8.2	10.8	12.3	43.2	13.4			31.6	11.3
Bismarck	5.4	7.2	5.5	4.1	6.0	4.6	6.0	4.1	5.7	5.0
Brakke	58.0	12.9	69.0	13.4	75.0	16.2	56.0	11.7	64.5	13.6
Buffalo	14.6	18.6	5.4	15.4	16.8	18.6	10.2	12.3	11.8	16.2
Burke	15.0	16.4	15.0	17.4	13.0	16.6	15.0	14.8	14.5	16.3
Byron	384.0	27.0	468.0	16.3	936.0	48.0	420.0	26.4	552.0	29.4
Carthage	14.0	15.0	16.5	18.6	16.0	19.6	13.0	18.3	14.9	17.9
Catron Pond	78.0	8.4	17.5	11.8	150.0	8.7	108.0	3.0	88.4	8.0
Cottonwood	588.0	60.0	576.0	16.4	768.0	90.0	456.0	36.0	597.0	50.6
Durkee	31.2	8.2	7.0	3.4	43.2	9.0	33.6	8.2	28.8	7.2
Elm	64.8	16.3	72.0	15.6	180.0	19.2	15.6	11.4	83.1	15.6
Enemy Swim	12.0	13.2	7.2	12.0	10.8	12.1	8.4	8.4	9.6	11.4
Gardner	110.0	5.0	22.6	7.2	192.0	6.9	21.0	1.3	86.4	5.1
Hayes	35.0	8.2	37.0	9.0	42.0	9.7	35.0	9.0	37.3	9.0
Henry	105.0	16.2	77.0	15.0	109.0	15.6	54.0	86.3	86.3	15.6
Hendricks	10.0	16.0	4.8	13.0	10.2	14.6			8.3	14.5
Herman	35.0	19.5	53.0	18.7	66.0	21.1	46.5	11.9	50.1	17.8
Hiddenwood	36.3	15.0	11.4	14.8	58.8	19.2	15.6	5.8	26.6	13.7
Horse Thief	6.0	2.6	4.0	2.6	3.7	2.8	5.0	2.2	4.7	2.6
Iron Creek	2.4	1.0	4.0	0.7	3.0	1.4	3.0	1.0	3.1	1.0
Isabell	48.0	9.0	14.4	8.7	66.0	9.3	10.8	1.4	34.8	7.1

Table 11. (continued)

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium	So- dium	Pot- tas- sium
Kadoka	74.0	7.8	75.0	7.4	96.5	7.2	87.0	7.3	83.1	7.4
Madison	109.2	22.6	115.0	22.0	130.0	29.5	96.5	18.3	112.7	23.1
Merindahl	23.0	16.2	24.5	15.8	31.0	19.7	28.0	15.6	26.6	16.8
Mission	16.6	16.2	14.0	14.8	17.0	11.4	13.0	9.5	15.2	13.0
Mitchell	64.2	16.1	25.0	17.0	83.0	17.3	72.0	14.7	73.6	16.3
Newell	55.2	4.9	84.0	7.2	76.2	6.3	28.8	4.9	61.0	5.8
Pactola	4.8	1.3	3.5	2.8	3.3	3.6	3.0	2.9	3.7	2.7
Pocasse	66.0	17.2	14.4	16.5	132.0	19.0	7.2	14.4	54.9	16.8
Poinsett	67.2	22.4	15.6	22.2	74.4	24.0	60.0	19.4	54.3	22.0
Punished Woman	7.2	5.4	4.8	7.2	17.2	17.0	15.6	7.8	11.2	9.4
Rose Hill	14.5	12.0	15.0	12.2	16.5	13.4	11.0	11.3	14.3	12.2
Roy	52.8	70.8	13.2	40.2	50.4	36.6	37.2	30.6	38.4	44.6
Shadehill	204.0	8.6	37.8	11.2	336.0	7.2	84.0	4.5	165.5	7.9
Sheridan	6.0	8.4	7.5	16.9	6.3	5.4	7.2	5.2	6.8	9.0
Stephen Mission	27.6	13.2	27.0	12.0	35.0	12.6	27.0	11.2	29.2	12.3
Traverse	27.6	11.0	14.4	14.9	64.8	13.9			35.6	13.3
Twin (Sanborn)	380.0	48.0	365.0	48.0	340.0	43.0	500.0	52.8	396.3	48.0
Twin (Spink)	504.0	45.0	288.0	18.0	888.0	87.6	552.0	54.2	558.0	51.2
Vermillion	16.2	13.8	15.0	13.8	12.0	12.0	9.2	12.3	13.1	13.0
Waggoner	31.2	14.4	37.0	8.0	33.5	8.6	37.0	7.9	34.7	9.7
White	33.6	16.8	10.8	15.9	38.4	16.2			27.6	16.3

Sodium concentrations were lowest in the Black Hills division which had a regional mean of 5 mg/l. The highest concentration measured during this study was 597 mg/l in Cottonwood Lake, a semi-closed lake in Sully County. Relationships between physical divisions of the state or soil types and the concentration of sodium ions were not well defined.

Potassium concentrations were lower west of the Missouri River than east of the river. Mean concentration for all lakes west of the river ranged from 1-16 mg/l, with a combined average of 8 mg/l. Lakes east of the Missouri which occupied closed basins or were underlain by aquifers ranged from 18-51 mg/l, while the remaining lakes in eastern South Dakota ranged from 9-17 mg/l potassium ion.

#### Nitrogen and Phosphorous

Nitrogen and phosphorous are essential to living organisms due to their role in protein metabolism and energy transfer. Of all the elements present in living organisms phosphorous is likely to be the most important ecologically; it is more likely to limit production than any other element. Sawyer (1954) stated that nitrogen is often considered the element which determines the extent of biological productivity, and phosphorous is often considered the governor of the "biological machine" which controls the speed changes which can occur.

Lakes which receive sewage or much agricultural drainage have higher concentrations of total phosphorous than are found in

uncontaminated waters. Sawyer (1944) concluded that any lake showing concentrations in excess of 0.01 ppm of inorganic phosphorous and 0.30 ppm of inorganic nitrogen at the time of spring turnover could be expected to produce algal blooms of such density as to cause nuisance.

Chu (1942) investigated the requirements of algae. Using pure chemicals he found that phosphorous concentrations below 0.05 ppm limited growth. He also found that the nitrogen and phosphorous requirements for different planktonic algae were approximately the same.

The total fish production of Minnesota lakes was shown to rise linearly from 100 to 400 pounds per acre as the phosphorous increased from 0.2 to 1.4 mg/l (Moyle, 1956).

Lackey (1945) has related the production of algal blooms in Wisconsin lakes to the concentrations of nutrients in the water and hence to the amount of sewage entering the lake. Lackey and Sawyer (1945) stated that some of these lakes receive annual dosages of saline nitrogen as high as 422 pounds per acre.

Since nitrogen can be fixed from the atmosphere, it may be concluded that phosphorous is the key element in the fertilization of natural bodies of water. Therefore, control of the amount of phosphate entering lakes is critical to production. Srairnh and Pillai (1966) stated that it is possible to eliminate phosphorous from sewage by the use of chemical coagulents, but this treatment may be expensive and the stability of the resultant effluent may

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not be satisfactory. They suggest the most efficient biological treatment method, from other considerations, followed by tertiary treatment to economically remove phosphorous from sewage effluent.

Combined nitrogen in water exists as inorganic compounds (ammonia, nitrate, and nitrite) and organic compounds. In the present investigation ammonium nitrogen was determined by the direct Nesslerization method while concentrations of nitrate and nitrite nitrogen were determined by the Hach Nitra Ver II method.

Seasonal concentrations of nitrogen as ammonia and nitrate are shown in Table 12. Nitrogen is discussed as the total of ammonia and nitrate since both are available to plants. Concentrations of nitrogen did not show any consistent seasonal patterns.

Mean concentrations of nitrogen ranged from 0.34 to 3.07 mg/l. Lowest concentrations of nitrogen (less than 0.5 mg/l) were in lakes of the Black Hills which receive little agricultural or sewage pollution. These lakes include Horse Thief, Pactola, and Sheridan with 0.46, 0.34, and 0.50 mg/l of nitrogen, respectively. Lakes Amsden and Enemy Swim in the Coteau des Prairies also had low nitrogen concentrations with 0.36 and 0.48 mg/l, respectively.

Phosphorous is known to occur in several forms. Those of greater concern in natural waters are soluble phosphate phosphorous, soluble organic phosphorous, and particulate organic phosphorous (Reid, 1961).

During this study samples for phosphate determination were frozen prior to analysis. Other investigations have shown that freezing samples as a method of preservation changes the level of

Table 12. Average concentrations of ammonia and nitrate (as mg/l nitrogen) from lakes studied, July 1965 to July 1967.

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate
Alvin	1.43	0.4		0.2	0.87	0.6	0.79	0.6	1.03	0.4
Amsden	0.45	0.0	0.00	0.2	0.34	0.2			0.26	0.1
Angostura	0.18	0.2		0.0	0.93	0.5	0.21	0.2	0.44	0.2
Bigstone	0.00	0.1	0.27	0.8	0.24	0.7			0.17	0.5
Bismarck	0.41	0.7	0.88	0.1	1.78	0.4	0.49	0.5	0.89	0.4
Brakke	0.60	0.3	0.06	0.2	0.22	0.6	0.25	0.3	0.28	0.4
Buffalo	0.70	0.0	0.46	0.6	0.00	0.4	0.30	0.4	0.36	0.4
Burke		0.2	0.34	0.8	0.67	0.8	0.24	0.2	0.42	0.5
Byron	2.35	0.6	2.43	1.8	0.30	2.1	0.25	0.8	1.33	1.3
Carthage		0.2	0.20	0.6	0.29	0.3	0.78	0.6	0.42	0.4
Catron Pond	0.28	0.8	0.64	0.6	0.60	1.0	0.25	0.4	0.44	0.7
Cottonwood	0.29	0.4	0.98	0.8	0.36	1.1	0.25	0.3	0.47	0.6
Durkee	0.87	0.7	0.66	0.3	0.23	0.2	0.30	0.2	0.52	0.4
Elm	1.33	0.4	1.80	1.3	0.29	0.7	0.35		0.94	0.8
Enemy Swim	0.61	0.0	0.25	0.4	0.00	0.1	0.25	0.2	0.28	0.2
Gardner	0.00	0.8	0.49	0.6	0.37	0.5	0.00		0.22	0.6
Hayes	0.60	0.3		0.1	0.16	0.2	0.09	0.5	0.28	0.3
Henry		0.5	2.69	0.7	1.50	0.6	0.50	0.2	1.56	0.5
Hendricks	1.64	0.0		0.9	0.28	0.8			0.96	0.6
Herman	0.43	1.0		0.8	1.50	0.2	0.54	0.7	0.82	0.7
Hiddenwood	0.30	0.2	0.46	0.7	0.35	0.8	0.00		0.28	0.6
Horse Thief	0.00	0.4	0.20	0.1	0.70	0.4	0.12	0.0	0.26	0.2
Iron Creek	0.00	0.3	0.90	0.1	0.27	0.0	0.45	0.8	0.40	0.3
Isabell	0.05	0.4	0.60	0.5	0.36	1.1	0.25	0.1	0.32	0.5

Table 12. (continued)

Lake	<u>Summer</u>		<u>Fall</u>		<u>Winter</u>		<u>Spring</u>		<u>Mean</u>	
	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate	Ammo- nia	Ni- trate
Kadoka	0.40	0.8		0.4	0.55	0.5	0.32	0.6	0.42	0.6
Madison	0.41	0.8	1.28	0.6	0.52	0.8	0.62	0.5	0.71	0.8
Merindahl	1.32	0.7	0.99	0.1	0.40	0.4	0.09	0.2	0.70	0.4
Mission	0.40	0.8	0.16	0.2	1.35	0.4	0.22	0.4	0.53	0.4
Mitchell	0.00	0.4	1.12	0.2	0.24	0.1	0.24	0.4	0.40	0.3
Newell	0.74	0.2	2.43	0.6	0.14	0.6	0.19		0.88	0.4
Pactola	0.19	0.2	0.25	0.3	0.02	0.1	0.12	0.5	0.14	0.3
Pocasse	0.95	0.8	0.50	1.1	0.38	1.2	0.23		0.52	1.0
Poinsett	1.32	0.5	0.00	1.2	0.91	0.5	0.30	0.6	0.63	0.7
Punished Woman	0.00	0.4	0.32	0.6	0.25	0.5	0.30	0.1	0.22	0.4
Rose Hill		0.2	2.93	3.4	2.18	0.6	0.49	0.5	1.87	1.2
Roy	0.73	0.0	1.37	0.5	0.00	0.4	0.21	0.2	0.58	0.3
Shadehill	0.00	0.8	0.44	0.3	0.27	0.5	0.25	0.2	0.24	0.4
Sheridan	0.22	0.5		0.0	0.17	0.3	0.21	0.5	0.20	0.3
Stephen Mission	0.00	0.4	0.21	0.6	0.19	0.2	0.56	0.6	0.24	0.4
Traverse	0.68	0.0	0.25	0.7	0.31	0.4			0.41	0.4
Twin (Sanborn)		0.6	2.40	0.6	1.30	0.3	1.16	0.5	1.62	0.5
Twin (Spink)	0.86	0.9	1.86	1.2	0.54	1.1	0.41	0.6	0.92	1.0
Vermillion	0.00	0.4		0.1	0.68	0.4	0.70	0.2	0.46	0.3
Waggoner	0.31	0.5	0.21	0.5	0.03	0.3	0.21	0.3	0.19	0.4
White	0.52	0.0	0.27	0.6	0.00	0.8			0.26	0.5

inorganic phosphate (Gilmartin, 1967), and affects release of phosphate from algae which may be in the sample (Fitzgerald and Faust, 1967). Since freezing causes the above mentioned changes, phosphorous was reported only as total phosphate.

Seasonal concentrations of total phosphate are shown in Table 13. Due to problems encountered during analysis, many of these values may be higher than actually occurred. Mean concentrations of total phosphate ranged from 0.16 to 3.00 mg/l.

### Iron and Manganese

Iron and manganese may be highly important in certain lakes since most lakes contain relatively small amounts of these elements. Both iron and manganese form the active site of many enzymes and are involved in the photosynthetic process. Manganese is also essential for nitrate assimilation.

Mean concentrations of total iron (Table 14) ranged from 0.02 to 0.78 mg/l, but only five of the lakes studied had mean concentrations over 0.2 mg/l. Mean concentrations of manganese (Table 14) ranged from 0.02 to 0.55 mg/l.

### Silica

Silicon is almost universally present in some form in all natural waters. It is a major nutrient for diatoms and may be a limiting factor since it is normally found in low concentrations. Mean silica content ranged from 1.1 to 28.5 mg/l (Table 14), but only 10 lakes had over 15.0 mg/l. Silica concentrations of lakes in

Table 13. Average total phosphate concentrations (mg/l) of lakes studied, July 1965 to July 1967.

Lake	Summer	Fall	Winter	Spring	Mean
Alvin	0.21	0.49	0.09	0.26	0.26
Amsden	*1.15	*0.04	*0.18		*0.46
Angostura	*0.36	*1.18	0.05	0.10	0.42
Bigstone	*0.10	*1.57	*0.22		*0.63
Bismarck	*0.70	0.07	0.07	0.41	0.31
Brakke	*0.18	0.15	0.05	0.17	0.14
Buffalo	*0.00	*0.29	*1.02	*0.46	*0.44
Burke	*2.22	1.28	1.48	0.78	1.44
Byron	*2.17	*1.91	*1.01	*1.01	*1.52
Carthage	*2.08	0.20	0.25	0.71	0.81
Catron Pond	*1.00	*0.40	*0.18	*0.23	*0.45
Cottonwood	*4.25	*2.08	*2.48	*3.20	*3.00
Durkee	*0.33	*0.23	*0.08	*0.06	*0.17
Elm	*2.66	*3.06	*0.80	*0.41	*1.83
Enemy Swim	*1.24	*2.28	*1.16	*0.60	*1.32
Gardner	*0.20	*0.27	*0.07	*0.11	*0.16
Hayes	*0.13	*0.28	0.13	0.26	0.20
Henry		0.45	0.12	0.05	0.21
Hendricks	*0.76	*0.12	*0.18		*0.35
Herman	*2.54	2.26	1.50	1.13	1.86
Hiddenwood	*2.60	*0.44	*0.53	*0.30	*0.97
Horse Thief	*0.87	0.05	0.05	0.14	0.28
Iron Creek	*0.58	0.05	0.26	0.29	0.30
Isabell	*0.50	*0.15	*0.09	*0.07	*0.20
Kadoka	*1.14	*0.78	0.12	0.20	0.56
Madison	*1.24	0.60	0.20	0.19	0.56
Merindahl	0.09	0.10	0.05	0.55	0.20
Mission	*0.76	0.05	0.06	0.18	0.26

Table 13. (continued)

Lake	Summer	Fall	Winter	Spring	Mean
Mitchell	*0.58	0.07	0.07	0.23	0.24
Newell	*0.22	*0.30	*0.09	*0.09	*0.18
Pactola	*0.86	0.06	0.05	0.23	0.30
Pocasse	*0.48	*0.05	*0.06	*0.10	*0.17
Poinsett	*1.48	*1.71	*1.74	*0.86	*1.45
Punished Woman	*0.16	*0.64	*0.17	*1.01	*0.50
Rose Hill	*2.06	0.05	0.08	1.04	0.81
Roy	*0.12	*0.22	*0.27	*0.29	*0.22
Shadehill	*0.49	*0.61	*0.14	*0.46	*0.42
Sheridan	*0.35	*0.63	0.05	0.18	0.30
Stephen Mission	*1.31	0.14	0.08	0.30	0.46
Traverse	*1.22	*1.62	*0.57		*1.14
Twin (Sanborn)	*1.22	0.14	0.08	0.18	0.41
Twin (Spink)	*2.46	*1.06	*0.39	*0.55	*1.12
Vermillion	*1.37	*2.12	0.16	0.34	1.00
Waggoner	0.19	0.09	0.05	0.35	0.17
White	*0.05	*2.53	*1.14		*1.24

\* Due to problems in analysis these values may be higher than actually occurred.

Table 14. Mean concentrations of iron, manganese, and silica (mg/l) of lakes studied, July 1965 to July 1967.

Lake	Iron	Manganese	Silica
Alvin	0.08	0.47	3.0
Amsden	0.05	0.23	8.2
Angostura	0.07	0.10	5.6
Bigstone	0.04	0.47	14.9
Bismarck	0.78	0.22	7.1
Brakke	0.08	0.10	2.4
Buffalo	0.06	0.30	7.6
Burke	0.14	0.06	19.4
Byron	0.07	0.31	25.3
Carthage	0.10	0.36	11.9
Catron Pond	0.26	0.12	9.1
Cottonwood	0.11	0.02	20.8
Durkee	0.14	0.14	1.2
Elm	0.08	0.83	22.3
Enemy Swim	0.02	0.27	15.4
Gardner	0.10	0.03	1.1
Hayes	0.09	0.06	2.1
Henry	0.50	0.24	6.1
Hendricks	0.08	0.33	11.4
Herman	0.16	0.44	13.1
Hiddenwood	0.16	0.31	13.0
Horse Thief	0.17	0.15	4.0
Iron Creek	0.37	0.16	3.1
Isabell	0.18	0.08	1.3
Kadoka	0.14	0.24	2.8
Madison	0.08	0.15	4.9
Merindahl	0.06	0.11	4.5
Mission	0.06	0.12	15.6

Table 14. (continued)

Lake	Iron	Manganese	Silica
Mitchell	0.05	0.08	5.7
Newell	0.03	0.64	2.4
Pactola	0.04	0.10	4.4
Pocasse	0.17	0.46	19.5
Poinsett	0.15	0.32	19.9
Punished Woman	0.06	0.27	9.0
Rose Hill	0.11	0.30	2.4
Roy	0.06	0.12	4.1
Shadehill	0.06	0.10	3.7
Sheridan	0.19	0.08	2.6
Stephen Mission	0.09	0.18	6.7
Traverse	0.07	0.55	28.5
Twin (Sanborn)	0.30	0.28	9.6
Twin (Spink)	0.13	0.40	13.9
Vermillion	0.08	0.33	10.6
Waggoner	0.07	0.17	4.2
White	0.11	0.23	17.2

the unglaciated area west of the Missouri River ranged from 1.1 to 9.1 mg/l, while lakes east of the river had concentrations ranging from 2.4 to 28.5 mg/l.

Aluminum, Barium, Boron, Copper

These four elements were never found in high concentrations (Table 15). It is not known if any of these elements limit aquatic life in any of the lakes studied. Mean annual concentrations of these elements ranged as follow: aluminum, 0.00 to 0.04 mg/l; barium, 0.2 to 3.8 mg/l; boron, 0.06 to 1.60 mg/l; and copper, 0.00 to 0.33 mg/l.

Table 15. Mean concentrations of trace elements (mg/l) from lakes studied,  
July 1965 to July 1967.

Lake	Aluminum	Barium	Boron	Copper
Alvin	0.01	0.9	0.06	0.05
Amsden	0.03	0.7	0.13	0.20
Angostura	0.01	0.8	0.22	0.24
Bigstone	0.00	1.5	0.27	0.07
Bismarck	0.03	1.0	0.39	0.00
Brakke	0.00	0.5	0.20	0.02
Buffalo	0.01	1.6	0.24	0.08
Burke	0.02	0.9	0.12	0.03
Byron	0.03	1.9	1.60	0.05
Carthage	0.03	1.6	0.16	0.04
Catron Pond	0.04	2.4	0.05	0.09
Cottonwood	0.00	1.9	0.84	0.02
Durkee	0.00	1.1	0.12	0.00
Elm	0.00	0.3	0.36	0.01
Enemy Swim	0.02	1.0	0.19	0.19
Gardner	0.01	1.8	0.48	0.23
Hayes	0.01	0.8	0.18	0.03
Henry	0.00	1.1	0.45	0.05
Hendricks	0.01	2.5	0.28	0.33
Herman	0.02	3.8	0.22	0.02
Hiddenwood	0.02	2.5	0.14	0.03
Horse Thief	0.03	0.2	0.34	0.04
Iron Creek	0.00	0.3	0.50	0.00
Isabell	0.02	1.6	0.15	0.03
Kadoka	0.02	1.2	0.61	0.01
Madison	0.01	0.6	0.24	0.03
Merindahl	0.00	0.8	0.63	0.10
Mission	0.02	0.6	0.05	0.04

Table 15. (continued)

Lake	Aluminum	Barium	Boron	Copper
Mitchell	0.01	1.2	0.49	0.02
Newell	0.01	0.2	0.16	0.02
Pactola	0.02	0.6	0.38	0.03
Pocasse	0.01	2.6	0.14	0.06
Poinsett	0.01	0.5	0.14	0.04
Punished Woman	0.01	2.1	0.10	0.03
Rose Hill	0.02	1.0	0.25	0.04
Roy	0.02	0.6	0.75	0.07
Shadehill	0.00	1.5	0.92	0.07
Sheridan	0.02	0.8	0.18	0.02
Stephen Mission	0.02	1.4	0.08	0.07
Traverse	0.04	2.2	0.32	0.08
Twin (Sanborn)	0.02	3.2	0.84	0.13
Twin (Spink)	0.03	2.9	1.23	0.03
Vermillion	0.02	1.1	0.11	0.06
Waggoner	0.01	0.8	0.15	0.04
White	0.02	1.0	0.12	0.18

## Biological Considerations

### Chlorophyll Concentrations

Concentrations of chlorophyll is commonly used as a measure of standing crop of phytoplankton. Since photosynthesis is dependent upon chlorophyll it would seem that primary production could be calculated from chlorophyll content. Steele and Baird (1961) and Ryther (1956) have pointed out many limitations of this method. The more important of these limitations are as follows: 1) chlorophyll measured in water is partly inactive, 2) assimilation numbers (mg carbon assimilated per mg chlorophyll a per hour) vary with light intensity, turbidity, and total dissolved solids, 3) there is a diurnal fluctuation in chlorophyll content, 4) the ratio of the chlorophyll components a, b, and c varies in different organisms, and 5) photosynthesis is light adaptive. Consequently, knowledge of chlorophyll content alone cannot provide a basis for productivity estimates. Data concerning chlorophyll a concentrations (Table 16) are used as general indices of standing crop of phytoplankton.

Chlorophyll a is the predominant chlorophyll pigment in both Chlorophyceae and Myxophyceae (Strickland, 1960). Chlorophyll b is present in Chlorophyceae but is a minor pigment. Since most dense algal blooms encountered were of Myxophyceae, and because the Richard-Thompson method does not give sufficient differentiation of pigments to enable taxonomic sorting (Strickland, 1960), only quantities of chlorophyll a were determined.

Table 16. Average concentrations of chlorophyll a (mg/l) from lakes studied, July 1965 to July 1967.

Lake	Summer	Fall	Winter	Spring
Alvin	4.393	10.584	0.245	1.233
Amsden	3.937	1.412		1.318
Angostura	.074	0.242	.065	.067
Bigstone	15.865	0.205		3.115
Bismarck	0.658	0.845	0.380	2.926
Brakke	2.456	1.032	0.202	0.099
Buffalo	1.178	0.733	1.554	1.094
Burke	3.993	0.886	0.378	0.389
Byron	6.260	1.114		0.072
Carthage	0.779	0.098	0.072	0.092
Catron Pond	19.114	0.366		0.219
Cottonwood	3.098	4.355		
Durkee	4.461	0.544		0.213
Elm	4.929	0.563		0.185
Enemy Swim	0.861		0.074	
Gardner	0.662	0.059		0.073
Hayes	0.346	0.188	1.669	0.138
Henry	0.464	0.069	0.034	0.149
Hendricks	7.079	1.207	0.700	4.670
Herman	14.256	0.603	0.070	5.686
Hiddenwood	3.268			0.500
Horse Thief	1.035	0.187	1.138	1.481
Iron Creek	0.834	0.288	0.190	0.416
Isabell	0.952		0.328	
Kadoka	0.328	0.378	0.161	0.098
Madison	16.638	0.366	9.149	2.648
Merindahl	0.446	1.967	0.152	0.810
Mission	1.266	0.247	0.077	1.810

Table 16. (continued)

Lake	Summer	Fall	Winter	Spring
Mitchell	0.870	0.504	0.456	0.983
Newell	0.352			
Pactola	0.175	0.134	0.113	0.176
Pocasse	18.037	0.658		2.341
Poinsett	13.811	5.073	3.565	1.667
Punished Woman	0.745	0.766	2.345	1.951
Rose Hill	1.079	2.077	0.850	1.820
Roy	1.765		0.230	
Shadehill	0.095			0.258
Sheridan	0.219	0.451	0.922	0.405
Stephen Mission	1.144	1.232	1.016	2.279
Traverse	7.804			1.937
Twin (Sanborn)	0.617	0.531	1.200	1.334
Twin (Spink)	4.122	3.308		1.226
Vermillion	2.560	9.120	0.044	1.788
Waggoner	9.662	0.563	0.280	0.280
White	8.746		1.205	2.523

Highest concentrations of chlorophyll a were observed in the late summer or fall and were usually caused by blue-green algae.

Aphanizomenon and Microcystis were the genera forming most of the nuisance algal blooms.

The highest seasonal concentrations of chlorophyll a are compared with the mean annual concentrations of total nitrogen and total phosphate in Table 17. Little association was observed between total phosphate ion concentration and concentration of chlorophyll a. This relationship may have been obscured because of problems encountered during analyses for phosphate. Lakes which had high mean total nitrogen content (more than 0.75 mg/l) often developed dense blooms of phytoplankton.

Table 17. Comparison of highest average chlorophyll a, mean total nitrogen, and mean total phosphate concentrations (mg/l) of lakes studied, July 1965 to July 1967.

Lake	Chlorophyll <u>a</u>	Nitrogen	Phosphate
Catron Pond	19.11	1.14	0.45
Pocasse	18.04	1.52	0.17
Madison	16.64	1.51	0.56
Bigstone	15.86	0.67	0.63
Herman	14.26	1.52	1.86
Poinsett	13.81	1.33	1.45
Alvin	10.50	1.43	0.26
Waggoner	9.66	0.59	0.17
Vermillion	9.12	0.76	1.00
White	8.74	0.76	1.24
Traverse	7.80	0.82	1.14
Hendricks	7.08	1.56	0.35
Byron	6.26	2.66	1.52
Elm	4.93	1.74	1.73
Durkee	4.46	0.92	0.17
Cottonwood	4.36	1.07	3.00
Twin (Spink)	4.12	1.92	1.12
Burke	3.99	0.92	1.44
Amsden	3.94	0.36	0.46
Hiddenwood	3.27	0.88	0.97
Bismarck	2.92	1.29	0.31
Brakke	2.46	0.68	0.14
Punished Woman	2.34	0.62	0.50
Stephen Mission	2.28	0.64	0.46
Rose Hill	2.08	3.07	0.81
Merindahl	1.97	1.10	0.20

Table 17. (continued)

Lake	Chlorophyll a	Nitrogen	Phosphate
Mission	1.81	0.93	0.26
Roy	1.76	0.88	0.22
Hayes	1.67	0.58	0.20
Horse Thief	1.48	0.46	0.28
Twin (Sanborn)	1.33	2.12	0.41
Buffalo	1.18	0.76	0.44
Mitchell	0.98	0.70	0.24
Isabell	0.95	0.82	0.20
Sheridan	0.92	0.50	0.30
Enemy Swim	0.86	0.48	1.32
Iron Creek	0.83	0.70	0.30
Carthage	0.78	0.82	0.81
Gardner	0.66	0.82	0.16
Henry	0.46	2.06	0.21
Kadoka	0.38	1.02	0.56
Newell	0.35	1.28	0.18
Shadehill	0.26	0.64	0.42
Angostura	0.24	0.64	0.42
Pactola	0.18	0.34	0.30

## SUMMARY

Maximum water temperature at the surface reached 28 C; temperatures of 23 to 25 C were common. Most of the lakes studied exhibited continuous circulation except when ice covered. Temperature in these lakes seldom varied more than 3 C from top to bottom at any given time. Six of the 45 lakes studied developed thermoclines. Maximum ice depth encountered ranged from over 100 cm in the northeast to about 20 cm in the southwest.

Light transmission was influenced by turbidity and varied greatly among lakes and within individual lakes during the year. The depth to which one per cent of the incident radiation penetrated the water varied from 0.2 m (Kadoka Lake) to 12.0 m (Pactola Lake). Turbidity was generally highest in the fall or summer due to stirring of sediments by wind action and development of algae blooms.

Dissolved oxygen concentrations ranged from near saturation to less than the recommended minimum for fish life. Many lakes showed decreased oxygen near the bottom. Lowest concentrations were encountered during the winter. All lakes studied were basic ranging in pH from 7.1 to 11.3. The pH was quite uniform throughout individual lakes or slightly more acid at the bottom.

Variations in specific conductance of lake waters were associated with physical divisions of the state and various drift types of the Wisconsin glacial age. Specific conductance of lakes occupying open basins was lowest in the unglaciated area west of the Missouri River and highest in the glacial drift east of the Missouri River. Open

lakes in the Black Hills had the lowest specific conductance (70-230 micromhos at 25 C) while lakes in the Mankato drift were highest (330-1260 micromhos at 25 C). Closed lakes or lakes overlying glacial drift aquifers had higher specific conductance than those lakes occupying open basins. Individual lakes showed maximum specific conductance during the winter due to concentration of the mineral elements under the ice.

Most of the major anions and cations followed concentration patterns similar to those described for specific conductance. Alkalinity of the lakes east of the Missouri River tended to vary according to the depth of carbonate leaching in the soil. Relationships between sodium concentrations and physical divisions or drift types were not obvious. Ranges of mean concentration for the major anions and cations in lakes with open basins are as follows: carbonate 38-260 mg/l (as  $\text{CaCO}_3$ ), sulfate 10-520 mg/l, calcium 7-364 mg/l, magnesium 3-241 mg/l, sodium 3-210 mg/l, and potassium 1-17 mg/l.

Mean concentration of total nitrogen ranged from 0.34 to 3.07 mg/l, and total phosphate ranged from 0.16 to 3.00 mg/l among lakes studied. Mean concentrations of total iron and manganese ranged from 0.02 to 0.78 and 0.02 to 0.83 mg/l respectively. Silica concentrations in lakes west of the Missouri River ranged from 1.1-9.1 mg/l; lakes east of the Missouri River had concentrations ranging from 2.4 to 28.5 mg/l. Aluminum, barium, boron, and copper were never found in high concentrations but were present in most lakes studied.

Chlorophyll a concentrations were usually highest in the summer or fall due to the blooms of the blue-green algae Aphanizomenon or Microcystis. Little association was observed between concentrations of total phosphate and chlorophyll a. This relationship may have been obscured because of problems encountered during the analyses for phosphate.

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## APPENDIX





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